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# A facile method for the preparation of Covalent Triazine Framework coated monoliths as catalyst support - applications in C1 catalysis

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**ABSTRACT.** A *quasi* Chemical Vapour Deposition method for the manufacturing of well-defined Covalent Triazine Framework (CTF) coatings on cordierite monoliths is reported. The resulting supported porous organic polymer is an excellent support for the immobilisation of two different homogeneous catalysts: 1) an Ir<sup>III</sup>Cp\*-based catalyst for the hydrogen production from

formic acid, and 2) a Pt<sup>II</sup>-based for the direct activation of methane *via* Periana chemistry. The immobilised catalysts display a much higher activity in comparison with the unsupported CTF operated in slurry because of improved mass transport. Our results demonstrate that CTF based catalysts can be further optimised by engineering at different length-scales.

## 1. Introduction

Covalent Triazine Frameworks (CTFs) are a class of porous polymers synthesised through the trimerisation of aromatic nitriles.<sup>1-2</sup> Over the past few years CTFs have been proposed for a wide range of applications.<sup>3-4</sup> For instance, due to their high nitrogen content, as candidates for CO<sub>2</sub> capture and separation,<sup>5-7</sup> while their fully conjugated nature offers new possibilities in electro-<sup>8-10</sup> and photocatalysis.<sup>11-12</sup> In addition, CTFs have shown to bear excellent properties as supports for single site organometallic catalysts. This application holds a large promise owing to the free *quasi*-bipyridine moieties present within the framework, allowing coordination of a large array of organometallic compounds.<sup>13-18</sup>

The high stability of CTFs is, to large extent, related to the harsh conditions under which these materials are synthesised (excess of ZnCl<sub>2</sub> as a trimerisation catalyst at 500°C). Such extreme synthesis conditions usually lead to an ill-defined CTF powder that consist of large agglomerates of up to tens of micrometers in size. Such poorly defined particles are obviously not ideal for application in catalysis, especially when targeting liquid-phase reactions in stirred tank reactors. On one hand, loss of material as a result of attrition of the agglomerates makes catalyst recycling cumbersome. On the other hand, large particles, even when resulting from the agglomeration of smaller units, may impose serious diffusion limitations and, therefore, poor catalyst utilisation. Moreover, the complex hydrodynamics of slurries in stirred tank reactors makes scale-up

difficult and can lead to problems like poor suspension of particles and a non-uniform distribution of catalyst. In our recent work, we addressed the first issue by formulating CTF/polymer spherical composites. The obtained spheres showed high stability and recyclability during several catalytic runs.<sup>14</sup> However, CTF's porosity could not be completely preserved upon formulation. Moreover, the use of polymers as binders limits the application of the resulting composites to reaction media in which the polymeric component would not dissolve. In the present work, we demonstrate that both issues can be addressed by CTF structuring, using a monolith as a support. Due to the unusual synthesis conditions, we have developed a new manufacturing protocol for the polymerisation of CTFs, directly on the surface of a cordierite monolith. This method, which can be described as a *quasi* Chemical Vapour Deposition, is a straightforward approach that renders a very uniform and thin CTF coating. These coatings are stable, easy to handle, and can be used in the same way as the parent material to coordinate metal complexes. A comparison between the performance of the monolith-supported CTF and powder-form CTF after incorporation of the catalysts (i.e. IrCp\* or Pt, *vide infra*) clearly demonstrates the superior performance of the former, attributed to the better mass transport properties of the thin layers coated on the monolith channels.

## **2. Experimental Section**

### **2.1. Micro-CTF synthesis.**

To synthesise micro-CTF, a glass ampoule was charged with 2,6-pyridinedicarbonitrile (0.124 g, 0.96 mmol) and anhydrous ZnCl<sub>2</sub> (0.664 g, 4.8 mmol) in a glovebox. The ampoule was flame sealed and the mixture was heated at 500°C for 48 h and then cooled to room temperature. The product was consecutively washed in 5M HCl at 100°C, in NH<sub>4</sub>OH at 60°C, in H<sub>2</sub>O at 100°C and then in THF at 60°C, each step overnight. The washing steps might seem excessive, but were, as

we found, required to remove  $\text{ZnCl}_2$ . 200 ml of each solvent was used for the washing procedure. Finally, the powder was dried in vacuum at  $180^\circ\text{C}$  overnight.

## 2.2. Meso-CTF synthesis

The synthesis procedure is identical to the one of micro-CTF with a different ampoule loading: 2,6-pyridinedicarbonitrile (0.041 g, 320  $\mu\text{mol}$ ), 4,4'-biphenyldicarbonitrile (0.131 g, 640  $\mu\text{mol}$ ), and anhydrous  $\text{ZnCl}_2$  (0.664 g, 4.8 mmol). The yield for both micro- and meso-CTF was 90%.

## 2.3. Monolith Coating

For the coating with CTF, 600 cpsi (cells per square inch) monoliths were used having the following characteristics: channel diameter 0.93 mm, wall thickness 0.11 mm, open frontal area 80%, wall porosity 0.33). The monolith pieces were cut down to a suitable size (around 1 cm long and 0.4 and 0.5 cm height and width, respectively).

In order to prepare the micro-CTF coated monolith, the monolith cut was dipped in a solution of 0.6 g 2,6-bipyridinedicarbonitrile and 1.7 g anhydrous  $\text{ZnCl}_2$  in 5 mL acetone. For the preparation of the meso-CTF coated monolith, a cordierite monolith piece was dipped in a solution of 0.6 g 2,6-bipyridinedicarbonitrile, 1.9 g 4,4-biphenyldicarbonitrile and 1.7 g anhydrous  $\text{ZnCl}_2$  in 5 ml acetone. In both cases the monolith was stirred overnight in the solution at room temperature. Following this step, the acetone solution wetted (impregnated) support was dried at  $60^\circ\text{C}$  overnight. Next, the monolith was placed in a glass ampoule, evacuated and flame sealed and subjected into an oven to the following temperature program: heating rate  $1^\circ\text{C min}^{-1}$  -  $500^\circ\text{C}$  for 10 h – cooling rate  $0.16^\circ\text{C min}^{-1}$ . The coated monolith was then washed with 50% HCl overnight at  $100^\circ\text{C}$ . Following this, it was subsequently washed with 50%  $\text{NH}_4\text{OH}$  at  $50^\circ\text{C}$  overnight, with water overnight at  $100^\circ\text{C}$ , and finally with tetrahydrofuran (THF) at  $50^\circ\text{C}$  overnight. The washed monolith was vacuum dried at  $180^\circ\text{C}$ . This washing procedure was

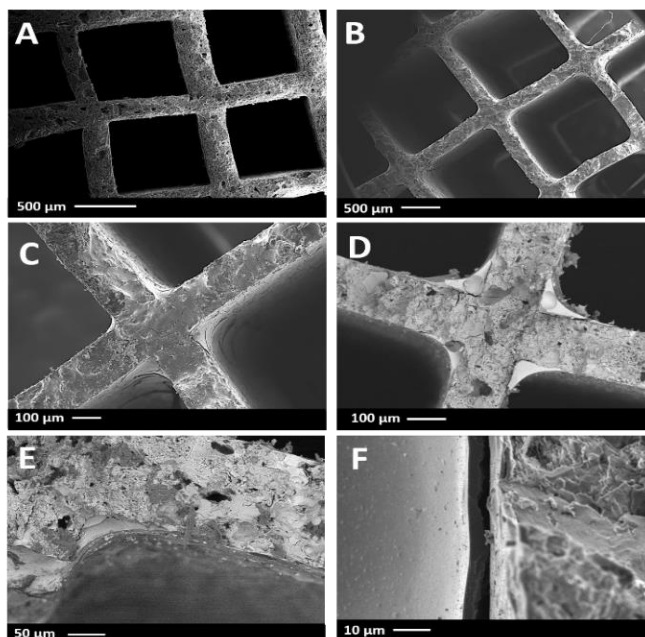
optimised in a previous work to remove as much Zn from the material as possible. The resulting CTF coatings are shown in Figure 1.

#### 2.4. Preparation of Ir@CTF@monolith

Firstly, a solution of 0.01 g  $[\text{IrCp}^*\text{Cl}_2]_2$  in 60 mL  $\text{H}_2\text{O}$  was prepared. Next, the CTF-coated monolith was stirred overnight in this solution at room temperature. Following this, the Ir exchanged monolith was washed in a mixture of 30 mL water, 30 mL DMF and 0.28 g trifluoromethanesulfonic acid (HOTf) at a temperature of 50 °C overnight in order to remove the  $\text{Cl}^-$  ions, and exchange the  $\text{Cl}^-$  ions by the weakly coordinated OTf. Subsequently, the monolith was washed with 50 ml water at 50 °C overnight. Finally, the Ir loaded monolith was vacuum dried at 60 °C overnight. Coordination of Ir to the powder was performed as described elsewhere<sup>13-14</sup> with an altered Ir loading (1 mg  $[\text{IrCl}_2\text{Cp}^*]_2$  and 100 mg micro- or meso-CTF powder).

#### 2.5. Preparation of Pt@CTF@monolith

Firstly, a solution of 0.01 g  $\text{K}_2[\text{PtCl}_4]$  in 10 mL  $\text{H}_2\text{O}$  was prepared. Next, the CTF-coated monolith was stirred overnight in this solution at 60°C. Following this, the Pt exchanged



monolith was washed in a 30 mL water. Finally, the sample was vacuum dried at 80 °C overnight. Coordination of Pt to the powder was performed as described elsewhere.<sup>15</sup>

**Figure 1.-** SEM micrographs of CTF@monolith. A: bare cordierite monolith, top view; B, C, D, E: CTF coated monolith, cross-sectional views; F: view on wall of a CTF coated monolith (*left*).

## 2.6. Formic Acid Dehydrogenation

The Ir@CTF@monolith catalyst samples with 0.045 mg Ir (0.23 mmol) were tested in a monolith stirrer reactor configuration. To this end, the structured support was attached to a mechanical stirrer and placed in a 50 mL aqueous solution of 3 M formic acid (0.15 mol). The vessel containing the reaction mixture was submerged in an oil bath and heated to the desired temperature of 80°C under continuous stirring. The reaction was carried out for 24 h. For recycling the catalyst, the coated monolith was washed with water at 50 °C and dried in vacuum at 60°C overnight. The Ir@CTF powder (50 mg) with 0.065 mg Ir (0.34 mmol) was tested as described elsewhere.<sup>13</sup>

## 2.7. Methane to Methanol Oxidation

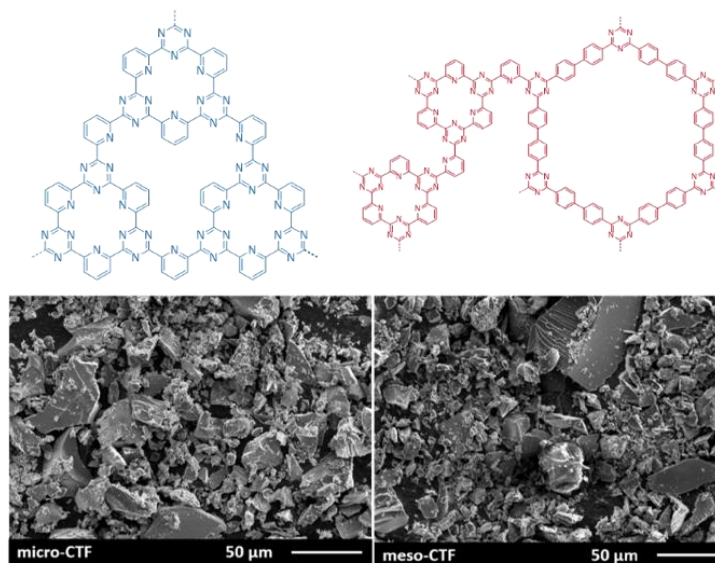
Catalytic tests were carried out in a 45 ml Titania grade 2 autoclave fitted with a Teflon insert in batch mode. The autoclave was filled with oleum (30% SO<sub>3</sub>, 10 mL) and the catalyst (CTF + monolith) was fitted inside the Teflon insert in a Teflon holder that provides a clamp to fit the monolith on top and a magnetic bar on the bottom (Figure S6). The autoclave was closed and flushed with helium. The reactor was pressurized to 40 bar CH<sub>4</sub> and heated to 180 °C at 2 °C min<sup>-1</sup>. The reaction time was 2 h, after which the autoclave was cooled down to room temperature. Depressurizing of the system was controlled by backpressure controller to avoid

foaming of the reaction medium. Once depressurized and helium flushed, the autoclave was opened and the monolith was taken out of the reaction medium and the liquids were collected in a separate glass vial for analysis quantification. Before any recycling experiment, the monolith was previously washed and dried at 80 °C under reduced pressure. Once dry the procedure followed for recycling was identical to that described above.

### **3. Results and Discussion**

In a previous work, when comparing the catalytic performance of Ir@CTF-polymer composite spheres in the dehydrogenation of formic acid, we observed a lower catalytic activity per Ir site (TOF) than that of the non-composite material.<sup>14</sup> Similar observations have been made for other heterogenised CTF supported catalysts<sup>15</sup> when comparing them to their homogeneous counterparts. This is usually attributed to diffusion limitations derived from heterogenisation. It is indeed straightforward to envisage that active site accessibility is easier for a dissolved homogeneous catalytic complex than that for a porous solid. In order to unravel differences in catalytic performance and to further improve the Ir-based catalyst utilisation, we have modified catalyst accessibility and mass transport at two different length scales. On the one hand, CTFs with two different pore architectures were synthesised: the one denoted as micro-CTF contains only micro-pores, while meso-CTF contains both micro- and meso-pores in a range from 1.5 to 10 nm.<sup>14, 19-20</sup> These materials differ in porosity and nitrogen content, while their chemical and thermal stability are alike. On the other hand, the performance of thin CTF coatings on a highly open support (*i.e.* a monolith) is compared with that of the bulk material (powder, Figure 2).





**Figure 2.-** top: Structures of micro-CTF (left) and meso-CTF (right); bottom: SEM micrograph of CTF powders (after crushing and reported washing procedure).

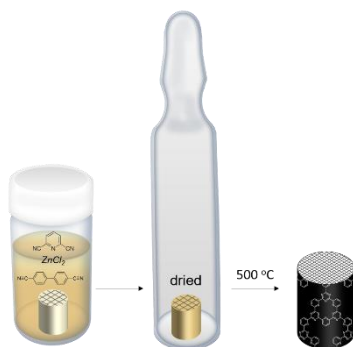
Monolithic structures are widely used in off-gas treatment, *e.g.* automotive exhaust gas purification<sup>21</sup> and de-NO<sub>x</sub> of power plant stack gases<sup>22</sup>. Monoliths are also considered attractive catalyst support alternatives compared to conventional carriers in other heterogeneous catalysis applications, including liquid-phase catalysis.<sup>23-26</sup> Monoliths feature a high void fraction and large geometric surface area, which results in a large contact area between the catalyst and reactants and low-pressure drop under flow conditions.<sup>24</sup> The most widely used monolith support material is based on cordierite ( $2\text{MgO}\cdot 2\text{Al}_2\text{O}_3\cdot 5\text{SiO}_2$ ) and it contains a macro-porous wall structure.<sup>27-29</sup> This macro-porous space may create an obstacle for applying a thin uniform layer of the catalyst, since it can lead to deposition of ineffective active phase within the macro-pores, where it is not accessible for substrates.<sup>30</sup> In this work, this property was partially turned to our advantage; the cordierite porosity was found to be crucial to achieve adhesion of the CTF coatings.

CTFs are typically synthesised in molten  $\text{ZnCl}_2$ , which serves both as a solvent and trimerisation catalyst. The synthesis occurs at elevated temperatures, at which the CTF precursors, except  $\text{ZnCl}_2$ , become volatile. Figure 3 summarises the synthesis protocol followed for the preparation of the supported CTFs: the cordierite monolith was impregnated with CTF precursors by filling the macro-porosity by using a low boiling point solvent (acetone). After acetone evaporation, the loaded monolith was placed in an ampoule and the polymerisation procedure was performed.

**Figure 3.-** Representation of the coating procedure steps

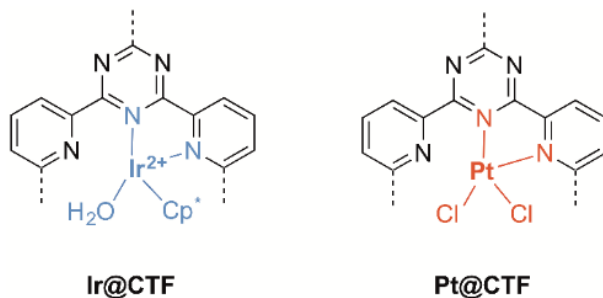
Since the macroporous monolith surface was homogeneously coated with zinc chloride, polymerisation of the evaporated monomers takes place selectively and right away on the surface of the support. The obtained films are in this way strongly attached to the cordierite surface (see Figure 1) and display excellent mechanical robustness during modification of the catalyst. Although the wall of the monoliths shows clearly a thin layer, thicker coatings are observed in the corners of the monolith channels. This relates to the capillary forces at the time of drying, which follows the acetone impregnation/exchange procedure inside of monolith.<sup>25</sup>

Additional experiments using a  $\gamma\text{-Al}_2\text{O}_3$  monolith were not successful because of dissolution of the monolith during the HCl post-treatment step to remove  $\text{ZnCl}_2$ . Although a coating could also be obtained using metal plates as dense supports, it was not mechanically rigid and was easily



removed from the surface when handling the material. Therefore, we conclude that the macroporosity of the selected cordierite monolithic support indeed plays an important role, „storing“ the precursors and  $\text{ZnCl}_2$  within its structure and, thus, serving in anchoring the CTF coating.

Since the proof of the pudding is in the eating, to demonstrate that the coatings observed in Figure 1 matter indeed the envisaged CTF, different characterisation techniques were applied. Thermo-gravimetric analysis in air (Figure S1) indicates a clear weight loss in the CTF@monolith sample at a temperature similar to that of the decomposition/oxidation of the CTF powder at around  $500^\circ\text{C}$ . From this analysis, a CTF coating of about 15 wt.% was calculated for the CTF@cordierite material. Having this in mind,  $\text{N}_2$  adsorption isotherms using the entire CTF@monolith piece were performed (nitrogen adsorption of bare cordierite monolith was found to be negligible). The obtained values were normalised to the mass of CTF in the CTF@monolith (Figure S2). The total loading of CTF per monolith is in the mg range. Therefore, the associated error to the measurement of the micro-CTF@monolith isotherms is expected to be somewhat high and thus it causes a mismatch when compared to the isotherm of the micro-CTF powder. In contrast, in the case of meso-CTF@monolith, the isotherms for the powder and the monolith are fairly identical. The nature of the nitrogen species from XPS



analysis was found to be identical in CTF and CTF@monolith (Figure S4).

**Table 1** TOFs of the hydrogen production from formic acid <sup>[a]</sup>

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The results described above present a compelling indication that the coating on the monolith bears the same nature as CTF powder. Therefore, the iridium catalyst active in hydrogen production from formic acid was anchored to the CTF-coating in a similar fashion as it was done

**Figure 4.-** Structures of the envisaged Ir@CTF (*left*) and Pt@CTF (*right*) catalysts.

to the powder or spheres (Figure 4).<sup>13-14</sup> Here, EDX analysis confirms that iridium was successfully coordinated to the bipyridinic moieties (Figure S4). The obtained catalyst is denoted as Ir@CTF@monolith. The performance of both micro- and meso-porous Ir@CTF@monoliths was tested in formic acid dehydrogenation and compared with that of powder samples with the same

metal loading (Table 1). All experiments were performed at identical conditions – 3M aqueous formic acid solution, 80°C, and similar stirring speeds. During recycling, the catalyst was stored under ambient conditions.

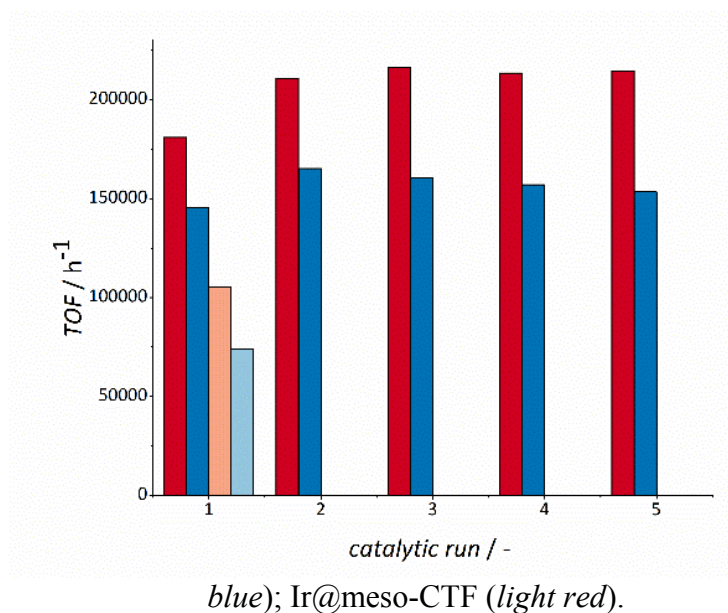
Entry	Sample	TOF [h <sup>-1</sup> ] <sup>[b]</sup>
1	Ir@micro-CTF@monolith <sup>[c]</sup>	156300 ± 160
2	Ir@meso-CTF@monolith <sup>[c]</sup>	207200 ± 180
3	Ir@micro-CTF <sup>[d]</sup>	73800 ± 110
4	Ir@meso-CTF <sup>[e]</sup>	105100 ± 125

<sup>[a]</sup> 50 ml 3M aqueous formic acid solution; <sup>[b]</sup> calculated from <sup>1</sup>H NMR analysis of formic acid; <sup>[c]</sup> 0.045 mg Ir (0.23 μmol); <sup>[d]</sup> 0.067 mg Ir (0.34 μmol); <sup>[e]</sup> 0.097 mg Ir (0.50 μmol).

From the catalytic results shown in Table 1, it is clear that porosity has an influence on catalyst performance. Comparison of the fully microporous version of the powder catalyst (entry 3) to its micro-mesoporous analogue (entry 4) showed that meso-porosity improved the catalytic activity by circa 30 %. While this activity trend was preserved when comparing the two monolith catalysts (entry 1 vs 2), it is obvious that by applying a coating onto a monolith an overall activity improvement of almost a 100 % (entry 1 vs 3 and 2 vs 4) is observed. We attribute these improvements to both the shorter diffusion path-length in the monolith coating (compare figure 2 (crushed powder) and figure 1 (coating on the monolith)) and to the improved mass transfer from the liquid to the coated channel wall. Nevertheless, application of catalyst performance test criteria indicates that the highly active Ir catalyst operates at the onset of mass transport interference.

Another clear advantage of the monolithic catalysts is that recycling and stability testing becomes straightforward. Figure 5 shows the observed performance for both monoliths over 5 consecutive catalytic runs. In both cases the first run displayed a slightly lower activity. We attribute this improvement to the occurrence of ligand exchange: most probably, the weakly coordinating triflate anion is replaced by formates in the first catalytic cycle. This was confirmed by EDX analysis, where fluorine is homogeneously distributed in the fresh catalyst and not detectable any more after several uses. Last but not least, it is worth highlighting that the *TOFs* obtained for the monolithic systems are the highest among those reported in the literature for heterogeneous systems for this reaction.<sup>16, 31-36</sup>

**Figure 5.-** *TOFs* of Ir catalyst for several catalytic runs of hydrogen production from formic acid: Ir@micro-CTF@monolith (*blue*); Ir@meso-CTF@monolith (*red*); Ir@micro-CTF (*light blue*); Ir@meso-CTF (*light red*).



Driven by these results and inspired by the pioneering work of Schüth *et al.*<sup>37</sup> on the immobilisation of the well-known Periana catalyst, we tested our monolithic supports in a more

challenging reaction: the selective partial oxidation of methane to methanol in oleum. Based on chemistry first developed by Shilov,<sup>38</sup> in the Catalytica/Periana System, a methane molecule is directly oxidised to methyl bisulfate when performing the reaction in oleum using a (bpym)PtCl<sub>2</sub> complex as catalyst.<sup>37</sup> The reaction proceeds at moderate temperatures (*ca.* 200 °C) and high methane pressure (*ca.* 40 bar). Recently, tetrachloroplatinate (K<sub>2</sub>[PtCl<sub>4</sub>]) was found to be active for the same reaction, with *TOFs* of 25000 h<sup>-1</sup>, more than 3 orders of magnitude higher activity than that for (bpym)PtCl<sub>2</sub>.<sup>39</sup> Years after the original work of Periana, Schüth *et al.* developed a heterogeneous counterpart by using a CTF as support (Figure 4) in a similar fashion to Ir, *vide supra*.<sup>15</sup> The Pt@CTF solid was found to be an active and recyclable catalyst, which is remarkable under the given reaction conditions. We, therefore, coordinated K<sub>2</sub>[PtCl<sub>4</sub>] to both micro- and meso-CTF coated monoliths. As in the case of the Ir complex, Pt was well dispersed over the shaped material (Figure S5). The monolithic pieces were attached within a Teflon holder (Figure S6) and submerged in fuming sulfuric acid (30% SO<sub>3</sub>) under 40 bar methane pressure. An unsupported version of the catalyst using microporous CTF with a similar Pt loading was also prepared and tested under slurry operation. Catalytic results are summarised in Table 2. While the Pt@micro-CTF@monolith preserved its integrity after the catalytic cycle, the mesoporous material containing monolith was seriously damaged under reaction conditions with a clear leaching to the oleum solution. We tentatively attributed the lower stability of the mesoporous coating to the higher accessibility, potentially leading to H<sub>2</sub>SO<sub>4</sub> intercalation and exfoliation under reaction conditions. Nevertheless, this sample attained a decent number of turnovers after 2 h (*TON*=2230). The Pt@micro-CTF@monolith appeared intact after 2 h under reaction conditions, attaining more than

**Table 2.** *TONs* of the production of methanol by methane oxidation<sup>[a]</sup>

Entry	Sample	Final MeSO <sub>4</sub> H concentration [M]	<i>TON</i> <sup>[b]</sup>
1	Pt@meso-CTF@monolith <sup>[c]</sup>	0.34	2230±25
2	Pt@micro-CTF@monolith <sup>[d]</sup>	0.64	5160±30
3	Pt@micro-CTF <sup>[e]</sup>	0.01	60±10

<sup>[a]</sup> 10 mL H<sub>2</sub>SO<sub>4</sub> (30% SO<sub>3</sub>), 40 bar CH<sub>4</sub>, 180 °C, 2h; <sup>[b]</sup> calculated from <sup>1</sup>H NMR analysis of methylbisulfate; <sup>[c]</sup> Pt content 0.3 mg; <sup>[d]</sup> Pt content 0.24 mg; <sup>[e]</sup> Pt content 0.33 mg.

5000 turnovers within 2 h. This activity is a factor two higher than that observed for its powdered counterpart. Once again, this demonstrates the advantages of catalyst structuring.<sup>28-29</sup> The obtained values showed that Pt@CTF@monoliths outperform the originally published Periana system with *TOFs* of 10<sup>-2</sup> s<sup>-1</sup><sup>37</sup> as well as Pt@CTF system of Schüth (*TON* of 426 after 2.5 h).<sup>15</sup> Although these values do not reach the highest reported *TOF* of *ca.* 25 000 h<sup>-1</sup> for this system reported in another work of Schüth under optimised conditions,<sup>39</sup> it opens up motivation for catalyst engineering at different lengthscales.

## 5. Conclusions

In summary, by applying a new synthetic protocol, CTF based stable films could be successfully deposited on the surface of cordierite monoliths. Extensive characterisation demonstrates that these coatings contain similar properties to those of the material in powder form. After coordination of an iridium complex, the obtained catalyst Ir@CTF@monolith was employed in the reaction of hydrogen production from formic acid. The obtained *TOF* is > 200 000 h<sup>-1</sup> in case of the Ir@meso-CTF@monolith. The catalyst was stable in ambient air and no



pre-treatment was required prior to utilisation. Pt@micro-CTF@monolith shows *TONs* of over 5000 after 2 h in the direct activation of methane. These results demonstrated that CTF based catalysts can be further engineered by optimisation at different length-scales (the micro-, meso-, and macro-level) alleviating internal and external mass transport limitations.<sup>28</sup>

#### ASSOCIATED CONTENT

##### **Supporting Information.**

Characterisation of CTF coating and the catalysts with TGA, Nitrogen physisorption, TGA, XPS, SEM

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##### **Author Contributions**

All authors have given approval to the final version of the manuscript. .

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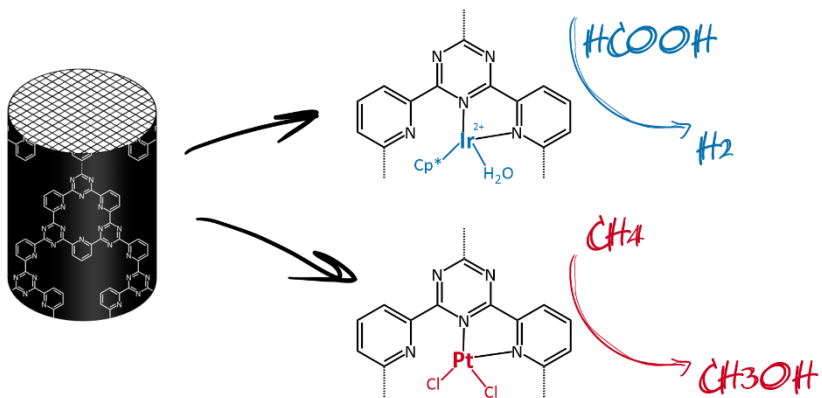
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