

Polarizable force fields for CO₂ and CH₄ adsorption in M-MOF-74

Becker, Tim M.; Heinen, J.; Dubbeldam, David; Lin, Li Chiang; Vlugt, Thijs J H

DOI

[10.1021/acs.jpcc.6b12052](https://doi.org/10.1021/acs.jpcc.6b12052)

Publication date

2017

Document Version

Final published version

Published in

The Journal of Physical Chemistry C

Citation (APA)

Becker, T. M., Heinen, J., Dubbeldam, D., Lin, L. C., & Vlugt, T. J. H. (2017). Polarizable force fields for CO₂ and CH₄ adsorption in M-MOF-74. *The Journal of Physical Chemistry C*, 121(8), 4659-4673. <https://doi.org/10.1021/acs.jpcc.6b12052>

Important note

To cite this publication, please use the final published version (if applicable). Please check the document version above.

Copyright

Other than for strictly personal use, it is not permitted to download, forward or distribute the text or part of it, without the consent of the author(s) and/or copyright holder(s), unless the work is under an open content license such as Creative Commons.

Takedown policy

Please contact us and provide details if you believe this document breaches copyrights. We will remove access to the work immediately and investigate your claim.

Polarizable Force Fields for CO₂ and CH₄ Adsorption in M-MOF-74

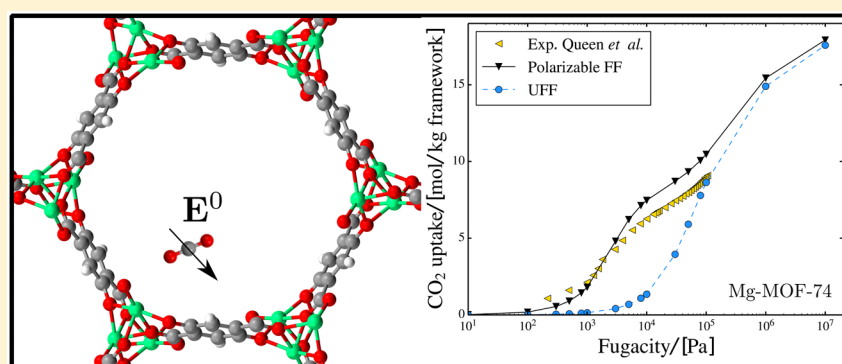
Tim M. Becker,[†] Jurn Heinen,[‡] David Dubbeldam,^{†,‡,§} Li-Chiang Lin,[§] and Thijs J. H. Vlugt^{*,†,§}

[†]Engineering Thermodynamics, Process & Energy Department, Faculty of Mechanical, Maritime and Materials Engineering, Delft University of Technology, Leeghwaterstraat 39, 2628CB Delft, The Netherlands

[‡]Van't Hoff Institute for Molecular Sciences, University of Amsterdam, Science Park 904, 1098XH Amsterdam, The Netherlands

[§]William G. Lowrie Department of Chemical and Biomolecular Engineering, The Ohio State University, 151 West Woodruff Avenue, Columbus, Ohio 43210, United States

Supporting Information



ABSTRACT: The family of M-MOF-74, with M = Co, Cr, Cu, Fe, Mg, Mn, Ni, Ti, V, and Zn, provides opportunities for numerous energy related gas separation applications. The pore structure of M-MOF-74 exhibits a high internal surface area and an exceptionally large adsorption capacity. The chemical environment of the adsorbate molecule in M-MOF-74 can be tuned by exchanging the metal ion incorporated in the structure. To optimize materials for a given separation process, insights into how the choice of the metal ion affects the interaction strength with adsorbate molecules and how to model these interactions are essential. Here, we quantitatively highlight the importance of polarization by comparing the proposed polarizable force field to orbital interaction energies from DFT calculations. Adsorption isotherms and heats of adsorption are computed for CO₂, CH₄, and their mixtures in M-MOF-74 with all 10 metal ions. The results are compared to experimental data, and to previous simulation results using nonpolarizable force fields derived from quantum mechanics. To the best of our knowledge, the developed polarizable force field is the only one so far trying to cover such a large set of possible metal ions. For the majority of metal ions, our simulations are in good agreement with experiments, demonstrating the effectiveness of our polarizable potential and the transferability of the adopted approach.

INTRODUCTION

The society's demand for energy and how it is currently satisfied interweaves strongly with anthropogenic CO₂ emissions and hence to the changing climate.^{1–3} It is evident that, to maintain present living standards, the energy sector needs to be altered drastically.⁴ New environmentally friendly ways of transforming energy have to be implemented on a large scale.⁵ This significant change of the energy sector is, however, still years from being fulfilled.⁶ New technologies need to be developed and further improved.⁷ Near-term measures include the considerable reduction of CO₂ emitted by conventional power plants.⁸ To reduce CO₂ emissions of current power plants, CO₂ needs to be separated from, e.g., the flue gas.^{8,9} Besides carbon capture, CO₂ removal is also crucial for other technologies, e.g., the purification of natural gas.^{10–12} A promising technology for the efficient separation of large quantities of CO₂ is the separation via solid adsorbents.^{13,14} In

this context, metal–organic frameworks (MOFs) have received substantial attention.^{6,15–17} MOFs are a relatively new and versatile type of material with various possible application areas such as in gas separation,^{18–22} gas storage,^{23–26} gas and liquid separation,^{7,27–31} catalysis,^{32,33} sensing,³⁴ drug delivery,^{35,36} microelectronics,^{37,38} and biotechnology.^{39–41} MOFs are constructed of metal ions or clusters connected by organic linkers.⁴² In recent years, a tremendous number of new MOFs has been synthesized⁴³ and an almost infinite number seems to be theoretically possible.⁴⁴ By adjusting the combination of metal ions and the organic parts, the properties of MOFs are widely tunable and materials with exceptionally large surface areas can be created.⁴⁵ The pore geometry can be customized

Received: November 30, 2016

Revised: January 25, 2017

Published: January 31, 2017

to enhance the separation of molecules due to the topology.^{46,47} In addition to the geometry, the chemical composition can be tuned to further improve separation performance.^{46,48,49} For instance, coordinatively unsaturated metal ions, so-called open-metal sites, can be embedded on the surface of the pore structure.⁵⁰ These metal ions are accessible for guest molecules and therefore interact strongly with certain adsorbate molecules.⁵¹ Therefore, the uptake of some adsorbate molecules can be increased significantly.⁵² Understanding and predicting the interaction of open-metal sites with adsorbates is crucial for the design of new customized adsorbent materials.⁵³ A challenge that is inherent with the enormous number of possibilities in the synthesis of MOFs is the selection of the best one for a particular application.⁴⁷ Experimental screening of hundreds of thousands of MOFs is impractical. A large effort has been made on developing computational screening approaches to facilitate the selection.^{13,44,54–56} Today, it is possible to predict adsorption properties for large sets of existing and hypothetical MOFs based on molecular simulations.^{57,58} A prerequisite for this kind of computational screening is a force field that represents the molecular interactions reasonably well for all materials under investigation. Unfortunately, the existing generic force fields do not fulfill this prerequisite for all MOFs.^{52,59–64} Especially, the promising class of MOFs with open-metal sites has been shown to be poorly described by generic force fields and research has been focused on developing improved force fields for these materials.^{62,65–68} To illustrate the failure of generic force fields, Figure 1 compares experimental measurements from Herm et

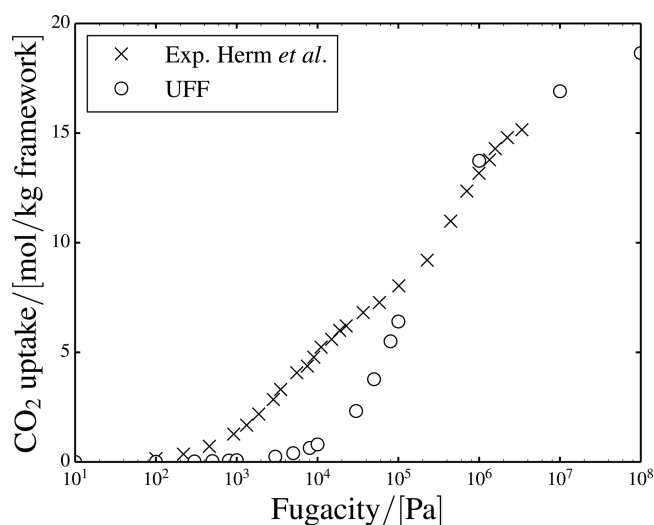


Figure 1. Comparison between the experimental adsorption isotherm of CO₂ in Mg-MOF-74 from Herm et al.¹⁸ and the simulated one using the UFF force field⁶⁹ for Mg-MOF-74 and the TraPPE force field⁷⁰ for CO₂ at 313 K.

al.¹⁸ to the CO₂ uptake in Mg-MOF-74 calculated from grand-canonical Monte Carlo simulations applying the generic UFF force field⁶⁹ for Mg-MOF-74 and the TraPPE force field⁷⁰ for CO₂ (i.e., standard generic force fields commonly used for porous materials^{64,71–73}).

The CO₂ uptake of Mg-MOF-74 predicted from molecular simulation for low fugacities with the UFF force field is considerably lower than the experimentally determined one. In particular, this region is highly relevant to carbon capture. The distinct inflection in the adsorption isotherm is also not

depicted, which suggests that the strong affinity of CO₂ close to the open-metal sites is not modeled correctly with the UFF force field. To obtain accurate force fields, several studies have been conducted in which force fields for individual MOFs are matched to interaction energies computed with quantum mechanical methods.^{49,62–64,66,74–76} In some of these studies, the applicability of the customized force fields was also investigated for MOFs with very similar topology and composition.^{49,63,64} Borycz et al.⁷¹ used this methodology to investigate the influence of the exchange of the metal ion for a MOF without open-metal sites. Addicoat et al.⁷⁷ designed an extension to the UFF force field to capture the structure of MOFs. Moreover, Vanduyffhuys et al.⁷⁸ developed a software package called QuickFF to automatically derive force fields for MOFs from ab initio input. The iso-structural M-MOF-74 has been pointed out to be a well suited study case to investigate the influence of different metal ions on the adsorption properties of small molecules^{51,64,79,80} and thereby further improve force fields. The pore structure of M-MOF-74 is only slightly influenced by the exchange of the metal ion, whereas the adsorption properties can change considerably. Understanding and describing the underlying interactions of open-metal sites with guest molecules is of fundamental interest⁶⁴ and can help to find trends and design even better adsorbent materials.⁷¹ M-MOF-74 is built of one-dimensional hexagonal pores with a diameter around 11 Å⁸¹ and exhibits a particular high density of open-metal sites.⁸² Figure 2 shows an extract from the periodic structure of Mg-MOF-74.

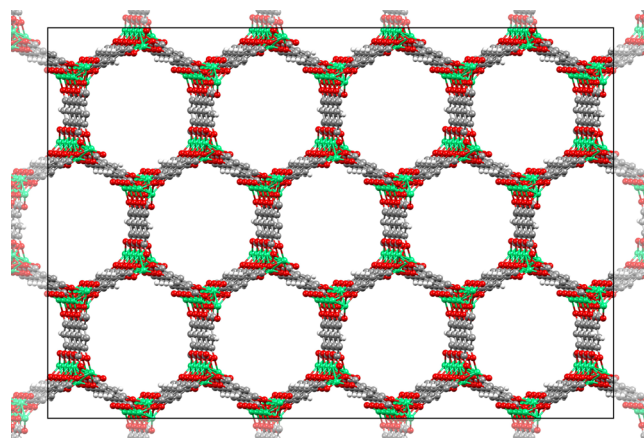


Figure 2. Extract from the periodic structure of Mg-MOF-74. Mg, C, O, and H atoms are represented in green, gray, red, and white, respectively.

In recent studies, Mg-MOF-74 has been shown to be a promising candidate for carbon capture due to its high CO₂ uptake capacity at low partial pressures^{18,19,52,81,83,84} and for natural gas sweetening.⁶⁴ Subsequently, M-MOF-74 has been extensively investigated for various gas separations.^{18,52,63,75,81,85} Among others, experimental studies include adsorption measurements of CO₂,^{18,19,46,51,52,81,83,84} CO,^{51,86} CH₄,^{51,79,87,88} C₂H₆,^{51,87–90} C₂H₄,^{87–90} C₂H₂,^{87,88} C₃H₈,^{87–90} C₃H₆,^{87–90} Ar,⁵¹ O₂,²⁰ and N₂.^{20,51} Adsorption sites have been investigated via neutron and X-ray powder diffraction to determine the binding geometry.^{46,79,91–93} As a complement to experiments, various quantum mechanical studies have been conducted to theoretically investigate adsorption sites,^{46,91} adsorption energies,^{79,80} and the underlying contributions and

mechanisms^{48,53} for a large number of guest molecules. Moreover, the mechanism of competitive adsorption has been studied by Tan et al.⁹⁴ These authors found that kinetic effects can play a significant role in the replacement of adsorbate molecules close to the open-metal sites. Molecular simulations have been used to investigate the adsorption behavior at uptakes larger than one guest molecule per open-metal site^{49,62,63,66,74–76} and the hopping of guest molecules between open-metal sites.^{95,96} Despite the significant progress, it is still a major challenge to accurately capture the change of interaction strength with varying metal ions in M-MOF-74 in molecular simulations. Several simulation studies have been conducted to reproduce the adsorption behavior of some of the M-MOF-74 structures.^{49,62,63,66,74–76} In these studies, standard interaction potentials are reparametrized to reproduce guest–host interactions from quantum mechanical calculations. In some quantum mechanical studies, it is suggested that guest molecules are polarized in the vicinity of the open-metal sites in M-MOF-74^{46,48,49,51,53,91} and that this interaction contributes to the enhanced CO₂ affinity. Standard force fields do not include this effect directly and therefore separate adjustments of the force field parameters may be necessary for every new structure. In this context, polarizable force fields for porous materials have been suggested,^{49,97–100} due to the observation of polarization in several other MOFs.^{101–103} However, we underline the potential of polarizable force fields for the study of adsorption in grand-canonical Monte Carlo simulations of M-MOF-74. We anticipate that considering polarization explicitly can help to create force fields that overcome the shortcomings of current generic force fields.

In this manuscript, we evaluate the potential of explicit polarization to improve the issue of limited force field transferability using MOFs with open-metal sites. In particular, we study the adsorption of CO₂ and CH₄ in M-MOF-74 with M = Co, Cr, Cu, Fe, Mg, Mn, Ni, Ti, V, and Zn. We extend the previously developed polarizable force field for CO₂ in Mg-MOF-74¹¹⁸ to structures based on nine more metal ions and CH₄ without additional fitting parameters. Subsequently, we conduct grand-canonical Monte Carlo simulations, and compare our results to results using other force fields^{64,69} and experiments.^{19,46,64,79,81,104} Thereby, it is shown that polarizable force fields have the potential to improve the transferability of force fields describing porous materials.

METHODOLOGY

Force fields describing intermolecular interactions are the foundation of molecular simulations.^{59,105} By definition, molecular simulations represent the behavior of a system for a given force field. The capability of the force field to describe the true molecular interactions determines its applicability.¹⁰⁶ An additional desirable characteristic of force fields is transferability.^{59,77,78} Ideally, a force field should be able to describe the experimentally observed behavior for a preferably large set of systems. Generic force fields like UFF,⁶⁹ DREIDING,¹⁰⁷ and OPLS¹⁰⁸ have been designed for organic, biological, and inorganic materials.⁵⁹ However, if the conditions of the system under investigation vary from the ones the force field was developed for, the resemblance of the real system behavior may be poor.⁵⁹ Many studies focus on regaining force field parameters to capture experimental behavior.^{49,62–64,66,74–76,109} This approach works well for individual systems.⁶⁴ Nevertheless, the transferability of the created force fields is likely to be limited to structures with very similar

topology and chemical composition.⁶⁴ In addition, for each pairwise interaction and all point charges, new parameters are required which are all mutually dependent.⁶⁴ Repeatedly readjusting force field parameters for every new system is cumbersome. The reason for the limited transferability of this approach could be attributed to the implicit consideration of the interactions that are exceptional for a particular system. Another disadvantage is that the predictive potential of molecular simulations is largely lost when force fields become completely empirical and need to be readjusted for every new system. A more sustainable approach is to develop force fields with a broader applicability due to a physically motivated extension which considers these exceptional interactions. Several studies focus on force field improvement of gas adsorption in MOFs.^{59,77,78} Unfortunately, a force field with general applicability for the adsorption of small molecules in MOFs does not exist.⁵⁹ Especially, the modeling of MOFs with open-metal sites represents a challenge.^{62,66,67} Quantum mechanical calculations of CO₂ adsorption in Mg-MOF-74 suggest that polarization of CO₂ in the vicinity of Mg ions is important and significantly contributes to the interaction energy.^{46,48,49,51,53,91} In contrast, charge transfer between CO₂ molecules and the MOF framework seems to be negligible.^{46,53,104} Several methods have been proposed for considering polarization in molecular simulations, i.e., the induced dipole method, the fluctuating charge method, and the shell method (also known as Drude oscillator and charge-on-a-spring model).^{110–115} For molecular dynamics simulations, these methods are well established.^{100,110,116,117} However, the many-body nature of polarization makes these algorithms more suitable for molecular dynamics simulation in which all molecules are moved in every simulation step. This is in contrast to Monte Carlo simulations in which usually only one molecule is moved.¹¹⁷ Hence, in Monte Carlo simulations, more steps are required to create independent configurations of the system. Normally, this is unproblematic, since the interactions need to be computed only for the moved molecule. However, when considering polarization, the interactions between all molecules change and have to be recomputed due to the many-body nature of polarization for every step. This leads to a less frequent consideration of polarization in Monte Carlo simulations. As described in our initial study,¹¹⁸ we use the procedure developed by Lachet et al.¹¹⁹ to mitigate this limitation. The procedure uses the induced dipole method in which the induction energy U_{ind} is expressed as

$$U_{\text{ind}} = -\frac{1}{2} \sum_{i=1}^N \boldsymbol{\mu}_i \cdot \mathbf{E}_i^0 \quad (1)$$

where $\boldsymbol{\mu}_i$ is the induced dipole, \mathbf{E}_i^0 is the permanent electric field created by the static partial charges at interaction site i , and N is the total number of interaction sites in the system. The energy contribution of the induction energy has to be computed in every Monte Carlo step. In this way, the difference in induction energy to the previous configuration can simply be added as another energy term in the acceptance rule of the Monte Carlo algorithm. Higher order induced multipoles are not explicitly incorporated in the induced dipole method. In a similar system, Lachet et al.¹¹⁹ estimated the related error to be less than 5% of the total induction energy. Special for the approach of Lachet et al.¹¹⁹ is that it accounts solely for polarization between the framework and adsorbate molecules and that it neglects

polarization caused by induced dipoles, so-called back-polarization. Using these assumptions, eq 1 can be rearranged to

$$U_{\text{ind}} = -\frac{1}{2} \sum_{i=1}^n \alpha_i \cdot |\mathbf{E}_i|^{0,2} \quad (2)$$

where α_i is the atomic polarizability of interaction site i and n is the number of interaction sites of the moved molecules. Thereby, an iterative scheme is avoided and the computational costs of the method are drastically reduced. In fact, the computational costs can be similar to simulations without considering explicit polarization. In the case of, e.g., a translation move of a single molecule only, the n interaction sites of this molecule have to be evaluated to determine the change in the induction energy. Lachet et al.¹¹⁹ showed that the error in energy introduced by this assumption is around 6% in a xylene NaY zeolite system. To verify the contribution of back-polarization in Mg-MOF-74, in Figure 3, the total interaction energy of a CO₂ molecule approaching the Mg ion with and without consideration of back-polarization for the developed polarizable force field is compared.

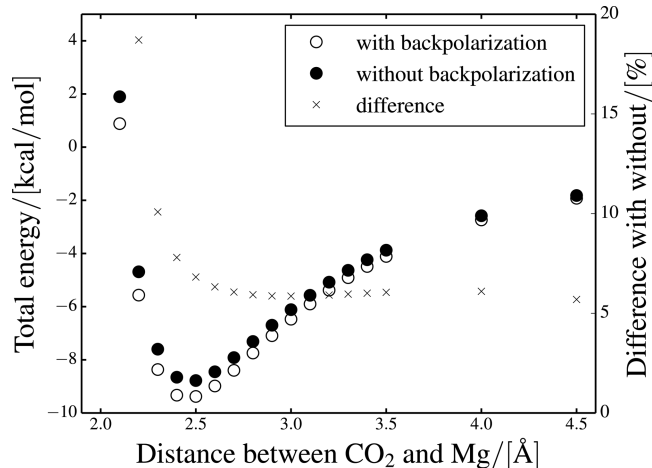


Figure 3. Total energy of a single CO₂ molecule in Mg-MOF-74 calculated using the developed polarizable force field as a function of the distance to the open-metal site. Comparison between interactions with and without back-polarization.

The influence of back-polarization increases with decreasing distance between the CO₂ molecule and the metal ion. For the most favorable position at approximately 2.4 Å, the difference in total energy is approximately 7%. This deviation seems to be acceptable in comparison with the considerable speedup of the simulations. Besides polarization, repulsion and dispersion interactions are considered via a standard Lennard-Jones potential and static charge distributions are modeled via point charges. When explicitly accounting for polarization, one has to ensure that the force field parameters describing the remaining interactions do not include an implicit polarization contribution which would have to be removed. Otherwise, the contribution of polarization would be double counted, once implicitly and once explicitly. The removal of implicit polarization is necessary if a standard force field is used as the starting point for the development of a polarizable force field, because current force fields are likely to be calibrated to reproduce certain experimentally observed properties. For example, the TraPPE force fields for CO₂ and N₂ are fitted to reproduce experimental vapor–liquid equilibria of the pure components and their

mixtures with alkanes without explicitly considering polarization.⁷⁰ Hence, in the fitting of these force fields, all present interactions are indirectly considered and the resulting potential parameters are effective parameters. As the starting point for our polarizable force field, we use the UFF⁶⁹ and TraPPE force fields.⁷⁰ These are standard force fields frequently used for molecular simulations of porous materials.^{64,67,71,120–123} To remove the contribution of implicitly considered polarization to the interaction potential, a global scaling parameter λ is applied to all Lennard-Jones energy parameters developed without explicit polarization. A simple procedure is chosen to verify the applicability of polarizable force fields rather than attempt to perfectly reproduce experimental results. Here, we reduce the Lennard-Jones energy parameters ϵ_i taken from the UFF and TraPPE force fields with respect to their atomic polarizabilities via

$$\epsilon_i^{\text{scaled}} = \epsilon_i \frac{(1 + \lambda) - \frac{\alpha_i}{\alpha_{\text{max}}}}{(1 + \lambda) - \frac{\alpha_i}{\alpha_{\text{max}}}\lambda} \quad (3)$$

where α_i and α_{max} are the atomic polarizabilities of interaction site i and the largest atomic polarizability, respectively. The scaling parameter λ can vary between 0 and 1. Thereby, it is assured that nonpolarizable interaction sites ($\alpha_i = 0$) are unchanged and that the potential energy parameters of the atoms with the largest polarizability are reduced the most. A more detailed description of the derivation and the algorithm of the presented method can be found in our previous publication.¹¹⁸ The required atomic polarizabilities α_i are taken from the literature.^{124,125} Many different values for atomic polarizabilities can be found for every atom.^{119,124–128} Their values can differ significantly depending on the experimental procedure or the theoretical assumptions made.¹²⁹ Hence, a global scaling factor ζ is used to adjust the magnitude of the atomic polarizabilities taken from the literature α_i^{lit} with respect to the chosen interaction potential according to

$$\alpha_i = \zeta \cdot \alpha_i^{\text{lit}} \quad (4)$$

Thereby, the ratio between the individual atomic polarizabilities is not affected to ensure a reasonable relative contribution of polarization between the atoms. This kind of scaling procedure for atomic polarizabilities is frequently used in the literature,^{130,131} and the scaled polarizabilities adopted in this study have comparable magnitudes to previous molecular simulation studies.^{99,119,132} In this manuscript, the values of ζ and λ are adjusted to reproduce the experimental adsorption isotherm for CO₂ in Mg-MOF-74. In a first step, the low fugacity region of the simulated adsorption isotherm and the heat of adsorption for CO₂ in Mg-MOF-74 are tuned by scaling all atomic polarizabilities with ζ . In the low fugacity region, CO₂ molecules adsorb close to the open-metal sites where polarization interactions are of particular importance. Subsequently, the scaling parameter λ is adapted to remove the implicit contribution of polarization from the Lennard-Jones potential. Therefore, the value of λ is lowered to match the high fugacity region of the experimentally determined adsorption isotherm. In this region, the centers of the channels of Mg-MOF-74 are filled with CO₂ molecules. The locations in the centers of the channels are further away from the open-metal sites, and therefore, polarization is less important. By applying this two-step procedure, we divide the interaction energy into

the underlying physical contributions without using an elaborated approach. For the remaining M-MOF-74 structures and for CH₄, the scaling factors determined for the Mg structure with CO₂ are used. Thereby, the transferability of the approach is investigated. The procedure is chosen to verify if the polarizable force field has the potential to describe the difference between the different metal ions embedded in M-MOF-74.

SIMULATION DETAILS

Grand-canonical Monte Carlo simulations implemented in the RASPA software package^{133,134} are conducted to compute the uptake and heats of adsorption of CH₄ and CO₂ in the different structures of the M-MOF-74 (M = Co, Cr, Cu, Fe, Mg, Mn, Ni, Ti, V, and Zn) family. The uptakes are computed for varying fugacities, for pure components and mixtures at 298 K. DFT-optimized, all atomic MOF structures with atomic charges are taken from Lee et al.⁸⁰ In the simulations, the structures are considered to be rigid. Lennard-Jones parameters for CH₄ and CO₂ are taken from the TraPPE force field.⁷⁰ Interactions between guest molecules are not modified and computed according to the TraPPE force field. The UFF force field⁶⁹ is used for the atoms of M-MOF-74. Cross-interactions are calculated via the Lorentz–Berthelot mixing rule from atomic parameters.¹³⁵ All Lennard-Jones energy parameters $\epsilon_i^{\text{scale}}$ used in the simulations are adjusted according to eq 3 with $\lambda = 0.7$. Thereby, we account for previously implicitly considered polarization. The Lennard-Jones potential is truncated at a cutoff distance of 12.8 Å without tail corrections. To mimic the behavior of the continuous system, i.e., a repetition of identical one-dimensional pores, periodic boundary conditions are applied in all directions (see Figure 2). The simulated system is composed of multiple unit cells to ensure a minimum distance of more than twice the cutoff radius between periodic images. The Ewald summation technique with a relative precision of 10⁻⁶ is used to calculate electrostatic interactions between static point charges.¹³⁶ Explicit polarization is considered via the induced dipole method.¹¹⁷ Polarization is exclusively considered between the framework and adsorbate molecules. Additionally, back-polarization is neglected to achieve reasonable simulation times. The required atomic polarizabilities α_i are taken from Shannon¹²⁴ and van Duijnen and Swart¹²⁵ and are scaled with a ζ value of 0.09. All force field parameters are summarized in Tables S1–S11 (Supporting Information). For the comparison with experimental results and simulation results of others, we use the Peng–Robinson equation of state to convert pressures to fugacities.¹³⁷ The DFT calculations to determine the orbital interaction energy (as explained below) are performed with the Amsterdam Density Functional (ADF) package.^{138,139} The B3LYP-D3 exchange–correlation functional^{140–144} is used with a TZP-STO basis set and a large frozen core. A fragment analysis is performed between CO₂ and Mg-MOF-74 to assess the net interaction between these two fragments. Using the energy decomposition analysis scheme by Ziegler and Rauk,^{145–147} the interaction energy ΔE_{int} between the two fragments is decomposed into

$$\Delta E_{\text{int}} = \Delta V_{\text{elstat}} + \Delta E_{\text{Pauli}} + \Delta E_{\text{oi}} \quad (5)$$

ΔV_{elstat} comprises classical electrostatic interactions between unperturbed charge distributions of the deformed fragments. ΔE_{Pauli} describes the Pauli repulsion energy and corresponds to the destabilizing interactions between occupied orbitals. The Pauli repulsion energy is responsible for steric repulsion. ΔE_{oi}

represents the orbital interaction energy which accounts for charge transfer and polarization.¹³⁸

RESULTS AND DISCUSSION

As an initial step in the evaluation of the developed polarizable force field, it is important to investigate the role of polarization in the adsorption of guest molecules in M-MOF-74. In Figure 4, we compare the polarization energy of a CO₂ molecule

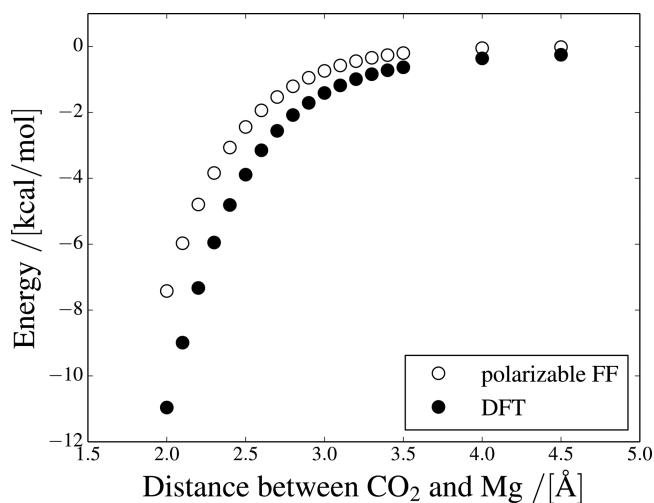


Figure 4. Comparison of the polarization energy computed with the developed polarizable force field without considering back-polarization and the orbital interaction energy from DFT calculations as a function of the distance between a CO₂ molecule and the Mg ion.

approaching the Mg ion of Mg-MOF-74 estimated with the developed polarizable force field to the orbital interaction energy calculated from ADF.

The orbital interaction energy should be a good approximation for the polarization energy, since no reaction is taking place and considerable charge transfer is not expected for very similar configurations of CO₂ inside Mg-MOF-74.^{46,53,104} For relevant distances, both methods show a comparable trend for the energy contributions. The most relevant distance between the CO₂ molecule and the Mg ion is where the total energy is the lowest (i.e., 2.3–2.5 Å, as shown in Figure 3). At this distance, the polarizable force field predicts that the polarization energy of a single CO₂ molecule in Mg-MOF-74 has a significant contribution of around 30% to the total energy. For larger distances, the contribution of polarization decreases rapidly. In our previous study,¹¹⁸ we investigated the quality of the developed polarizable force field by comparing the total energy of random CO₂ positions inside Mg-MOF-74 with detailed DFT calculations from Lin et al.⁷⁵ Thereby, also less favorable positions further away from the Mg ions are probed. These positions are occupied after all open-metal sites are saturated. In general, the polarizable force field describes most positions considerably better than the UFF force field and with a quality comparable to a nonpolarizable force field that has been developed by readjusting the majority of force field parameters.⁶⁴ The scheme applied here is considerably simpler.¹¹⁸ The resulting adsorption isotherms for CO₂ in Mg-MOF-74 in comparison to experimental measurements, the UFF force field, and the DFT-derived nonpolarizable force field of Mercado et al.⁶⁴ are shown in Figure 5a.

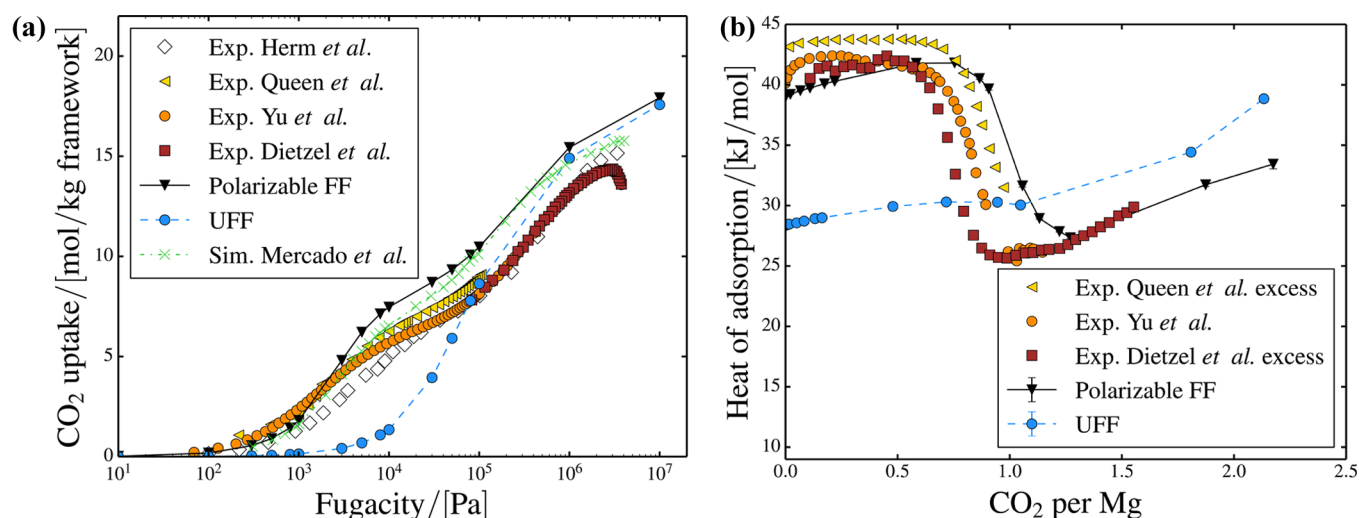


Figure 5. Comparison between the experimental results of Herm *et al.*¹⁸ (open), Queen *et al.*⁴⁶ (yellow), Yu *et al.*¹⁰⁴ (orange), and Dietzel *et al.*⁸¹ (brown) and simulation results using the developed polarizable force field (black), the UFF force field (blue) and the DFT-derived nonpolarizable force field of Mercado *et al.*⁶⁴ (green) for CO₂ in Mg-MOF-74. (a) Adsorption isotherm at 298 K (Herm *et al.*,¹⁸ 313 K); (b) heat of adsorption as a function of uptake.

The simulation results with the polarizable force field clearly display the inflection of the experimental adsorption isotherm. The predicted behavior is significantly better than that with the UFF force field. This is expected, because the scaling factors are adjusted to reproduce the experimental data. The overall agreement with the experimental measurements is comparable with the DFT-derived nonpolarizable force field of Mercado *et al.*⁶⁴ Both force fields can predict the low fugacity region which is particularly important for carbon capture. For higher fugacities, simulations with all compared force fields predict higher CO₂ uptakes in comparison to the experiments. As pointed out earlier, this can be attributed to the fact that a certain degree of inaccessibility due to diffusion limitation or defects in the crystal structure is inherent with experimental structures.^{49,81} In the limit of very high CO₂ uptakes, the guest–host interactions become less important and the adsorption is dominated by the accessible volume for CO₂.⁷¹ In the development of the used TraPPE force field, the CO₂–CO₂ interactions were adjusted to reproduce the vapor–liquid equilibria and it describes the density per void volume well. Therefore, the uptake of CO₂ predicted using the polarizable and UFF force field converges in the high fugacity region. It should be noted that Mercado *et al.*⁶⁴ scaled the calculated CO₂ uptakes with 0.85 to account for inaccessibility of open-metal sites. This scaling procedure mainly improves the agreement between experiments and computations for the high fugacity region. Figure 5b shows the heat of adsorption as a function of CO₂ molecules per metal ion. The distinct inflection of the adsorption isotherm caused by the strong affinity of the CO₂ molecule toward the metal ions is reflected by the change of the heat of adsorption with increasing gas uptake. The calculated heat of adsorption has an inflection at exactly one CO₂ per metal ion. Before and after the rapid decrease at one CO₂ molecule per metal ion, the heat of adsorption increases slightly. This increase can be related to a rise in the total number of adsorbed CO₂ and therefore a larger contribution of the CO₂–CO₂ interactions to the total energy. Similarly, the experimental heats of adsorption increase initially. In general, the experimental heats of adsorption have to be regarded with wariness. The heat of adsorption is not measured directly but

calculated according to $-q_{st}/R = \partial(\ln p)/\partial(1/T)$ at constant loading and averaged over adsorption isotherms at different temperatures.¹⁰⁴ The temperatures considered vary for all experimental studies. The experimental curves consistently show a drop in the heat of adsorption for lower ratios of the number of guest molecules and metal ions than the simulation results. Different sets of experimental adsorption isotherms show inflections at different uptakes of CO₂. This is another indication for defects in the crystal structure and the blocking of some of the metal ions of the experimental structures. Haldoupis *et al.*⁴⁹ further investigated the effect of blocking for the Co, Cu, Mn, and Ni based structures. These authors illustrate that varying levels of pore accessibilities can explain the discrepancy between different experimental studies. Especially, the unavailability of open-metal sites can explain the drop in the heat of adsorption prior to the complete saturation of these sites. This can be caused by residual solvent molecules binding to the open-metal sites. According to Haldoupis *et al.*,⁴⁹ these residual solvent molecules could reduce the number of accessible open-metal sites by 20–30%, while only slightly affecting the accessible surface area and the accessible volume. Previously, the good agreement between the experimental BET surface area and pore volume and the theoretical void space in the empty crystal structure made Dietzel *et al.*⁸¹ suggest that their MOF was fully activated. In contrast, Haldoupis *et al.*⁴⁹ concluded that a combined effect of crystalline defects and residual solvent molecules is most likely to cause the difference between simulations and experiments. The focus of this study is to evaluate the applicability of a polarizability force field for describing the interactions of guest molecules with different metal ions. In this regard, the consideration of residual solvent molecules and defects does not seem to be crucial. For the remaining M-MOF-74 structures, the parameters adjusted for Mg-MOF-74 and CO₂ are used. Thus, the calculated values are predictions based on the two global scaling parameters λ and ζ adjusted for Mg-MOF-74. To obtain an overview of the results, we divided the predictions for CO₂ in the M-MOF-74 structures into three groups. The first group consists of Co, Cu, Ni, and Zn, the second of Cr, Ti, and V, and the third of Fe and Mn. In Figure

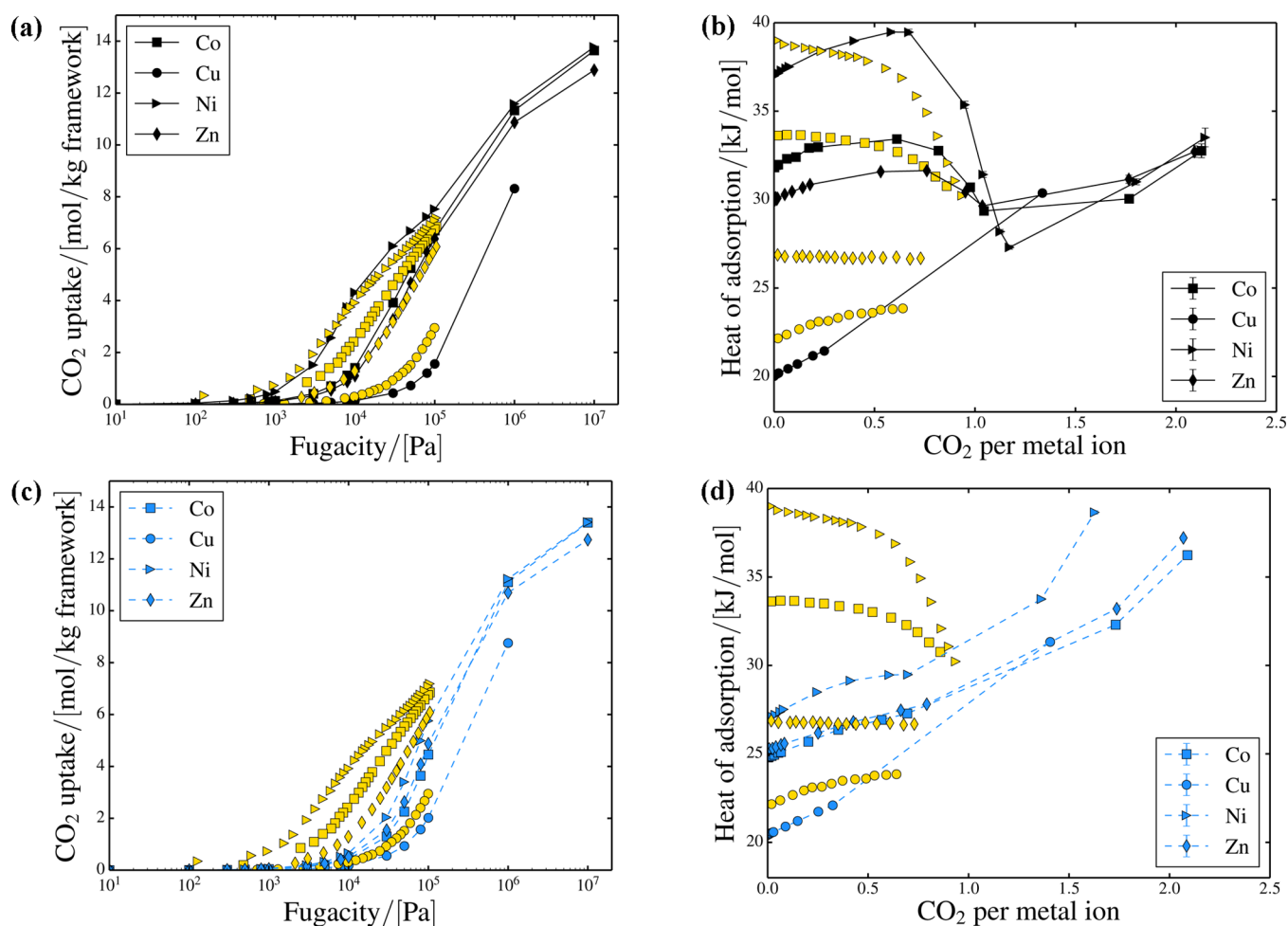


Figure 6. Comparison between the experimental results of Queen et al.⁴⁶ (yellow) and simulation results using the developed polarizable force field (black) and the UFF force field (blue) for CO₂ in the Co (■), Cu (●), Ni (▲), and Zn (◆) based structures. (a and c) Adsorption isotherms at 298 K; (b and d) heats of adsorption as a function of uptake.

6, the computational results for the first group are presented and compared to experimental measurements of Queen et al.⁴⁶ and the UFF force field.

For these structures, the developed polarizable force field is able to describe the experimental measurements well. This is most striking in comparison to the UFF force field which is not able to model the differences between the metal ions. For the polarizable force field, the largest deviations in the adsorption isotherms can be observed for Cu and Co based structures. In agreement with the experimental data, the simulation results for Co and Ni based MOFs show a less distinct inflection for the CO₂ adsorption isotherm than for Mg-MOF-74. The experimental adsorption isotherms for Cu and Zn based structures do not show an inflection for CO₂, which is also accurately predicted in the simulations applying the polarizable force field. The comparison to the experimental results of Yu et al.¹⁰⁴ and the simulation results of Mercado et al.⁶⁴ can be found in Figures S22–S11 (Supporting Information). The results of Mercado et al.⁶⁴ match the experimental adsorption isotherms well. It is worth mentioning that Mercado et al.⁶⁴ did not perform simulations for Cu-MOF-74. These authors did not develop a force field for Cu-MOF-74, because of an elongation of the unit cell in the *c*-direction⁸⁰ in comparison to the other M-MOF-74 structures. Although our results for the Cr based structure deviate from the experimental results

(compare Figure 6a), the elongation does not seem to be problematic for the general applicability of our approach. The calculated heats of adsorption shown in Figure 6b and d have a similar quality as that for the Mg based structure. The largest discrepancy between simulations and experiments can be observed for Zn-MOF-74. This is surprising, because the calculated adsorption isotherm agrees very well with experiments and is very similar to the calculated adsorption isotherm for Co-MOF-74 with a similar heat of adsorption. Similar to Mg-MOF-74, the heat of adsorption derived from experiments shows an inflection significantly before an uptake of one CO₂ molecule per metal ion. As mentioned previously, residual solvent molecules are likely to cause this shift in the heat of adsorption,⁴⁹ since less open-metal sites are accessible. The simulations predict a similar behavior for all structures after all open-metal sites are saturated with CO₂. In this region, the CO₂ molecules start to accumulate in the centers of the channels. The geometry of the channels is almost identical for all types of M-MOF-74, and the CO₂ molecules are sufficiently far away from the metal ions to not be significantly affected by polarization. Overall, the polarizable force field seems to have the potential to capture the different degrees of polarizations related to the different metal ions for these four structures. Moreover, in contrast to the UFF force field, the polarizable force field is able to predict the correct order of adsorption

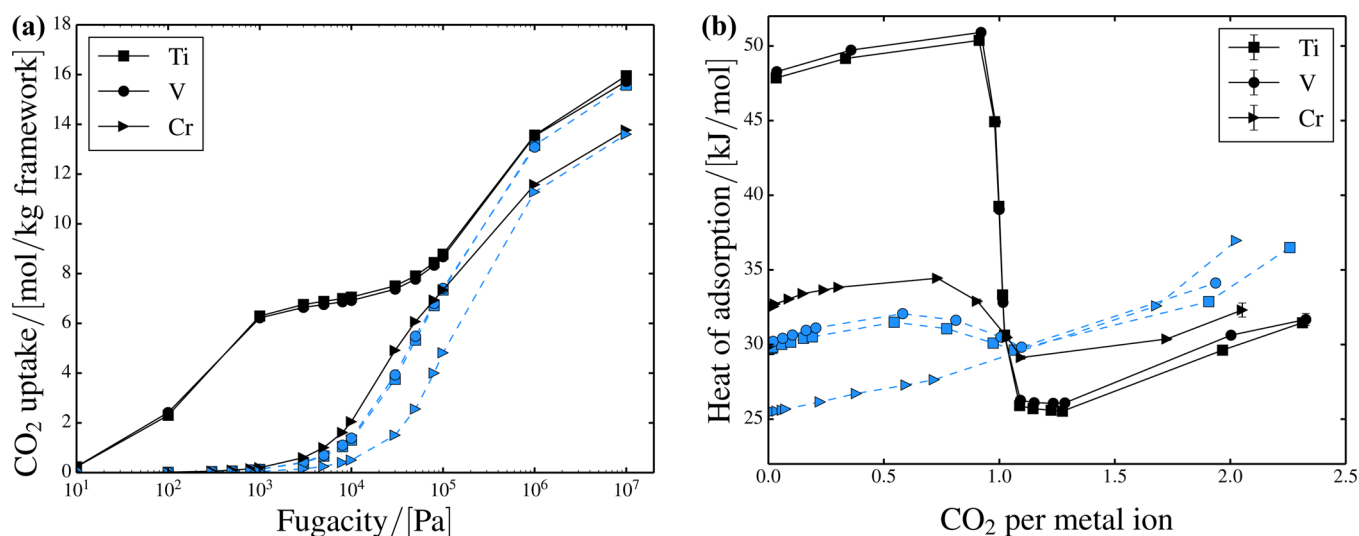


Figure 7. Comparison between the simulation results using the developed polarizable force field (black) and the UFF force field (blue) for CO₂ in the Ti (■), V (●), and Cr (►) based structures. (a) Adsorption isotherms at 298 K; (b) heats of adsorption as a function of uptake.

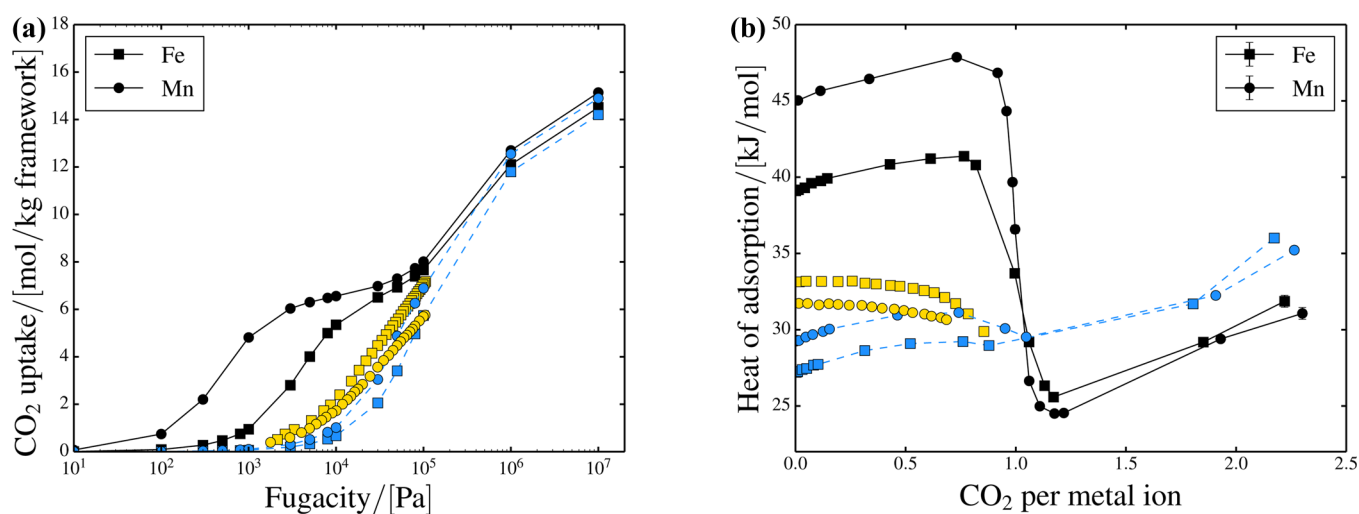


Figure 8. Comparison between the experimental results of Queen et al.⁴⁶ (yellow) and simulation results using the developed polarizable force field (black) and the UFF force field (blue) for CO₂ in the Fe (■) and Mn (●) based structures. (a) Adsorption isotherms at 298 K; (b) heats of adsorption as a function of uptake.

strength for the Co, Cu, Ni, and Zn based structures. The computationally predicted adsorption isotherms and heats of adsorption for the second group are compared to the computational results with the UFF force field in Figure 7.

For these structures, no experimental adsorption measurements are available. To the best of our knowledge, these structures belong to the group for which the experimental syntheses are still challenging.¹⁴⁸ The simulations predict a very distinct inflection for the adsorption isotherms of Ti- and V-MOF-74, while the one for Cr-MOF-74 does not show an inflection. The predictions for the Ti and V based structures agree with theoretical predictions of Park et al.⁴² These authors expect the structures to have even stronger interactions with CO₂ than Mg-MOF-74 which is supported by our simulations. The trend of the adsorption isotherms for the three MOFs is reflected in the heats of adsorption. The possibility to predict large differences between adsorption behavior shows further that polarizable force fields have the potential to describe such significant differences in the adsorption behavior. Again, the UFF force field predicts a totally different adsorption behavior

and a smaller difference between the metal ions (compare Figure 7). The remaining M-MOF-74 structures are based on Fe and Mn. A notably large discrepancy between simulations and experiments is found for these structures, as shown in Figure 8.

The experimental results for the Mn and Fe based structures are very similar. Both structures show weaker interactions between the metal ions and the CO₂ molecules than for the Mg based structure. As can be seen, the developed polarizable force field significantly overestimates these interactions. The UFF force field is able to capture the adsorption behavior better. Several reasons for the overestimation are possible, and we feel a combination of different effects is most likely. Interestingly, Mercado et al.⁶⁴ also failed to obtain a reasonable force field for Mn-MOF-74 based on fitting the interaction potential to quantum mechanical energies. This further suggests that the explanation for the failure could be rather complicated. For example, ferromagnetic effects which are not considered in the polarizable force field may play a more important role for Mn and Fe than for the other structures. Additionally, the initial

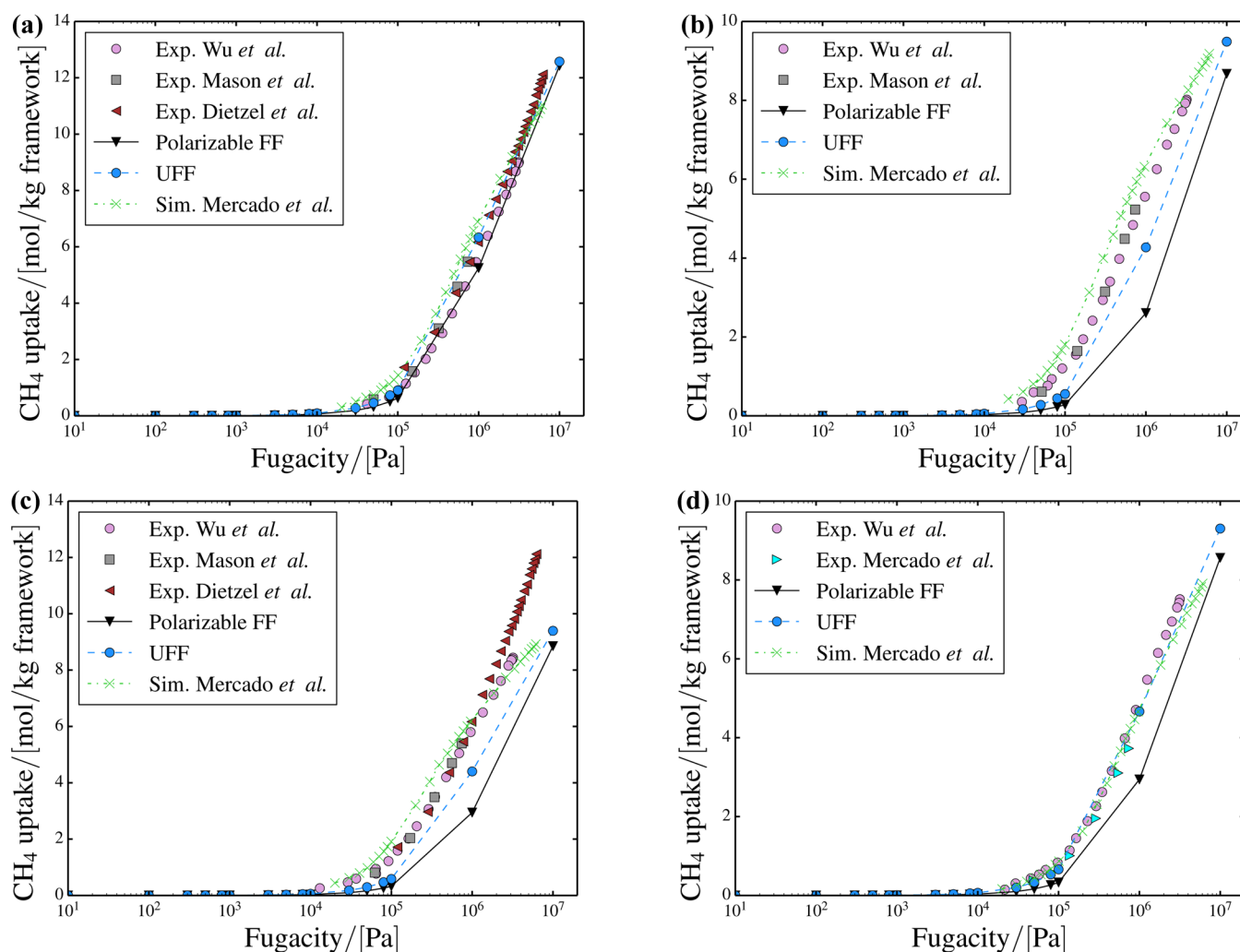


Figure 9. Comparison between the experimental results of Wu *et al.*⁷⁹ (violet), Mason *et al.*¹⁹ (gray), Dietzel *et al.*⁸¹ (brown), and Mercado *et al.*⁶⁴ (cyan) and simulation results using the developed polarizable force field (black), the UFF force field (blue), and the DFT-derived nonpolarizable force field of Mercado *et al.*⁶⁴ (green) for CH₄. (a) Mg-MOF-74, (b) Co-MOF-74, (c) Ni-MOF-74, (d) Zn-MOF-74.

force field parameters taken from the UFF force field for the two metal atoms could be of particularly bad quality. A comparison between the values for Fe from the DREIDING ($\epsilon/k_B = 27.677$ K, $\sigma = 4.045$ Å) and the UFF force field ($\epsilon/k_B = 6.542$ K, $\sigma = 2.59$ Å) shows the huge difference. Simulations based on the DREIDING parameters result in a totally different prediction of the adsorption behavior. This is illustrated by the Henry coefficients of CO₂ in Fe-MOF-74 in the limit of infinite dilution condition¹⁴⁹ we computed for both sets of force field parameters using Widom test particle insertions¹⁴⁹ (Supporting Information). The sensitivity of the system might also play an important role. In addition, the quality of the selected level of theory for the structure optimization could be better for some of the metal ions than for others.⁶⁴ The failure of the polarizable force field to predict the behavior of these two structures needs to be further investigated. Nevertheless, it does not diminish the potential of polarizable force fields for the description of MOFs with open-metal sites. To further verify the applicability of polarizable force fields, grand-canonical Monte Carlo simulations are performed for CH₄ in the M-MOF-74 series. The separation of CO₂ and CH₄ is industrially relevant.¹² Additionally, CH₄ is explicitly chosen to examine the suitability of the polarizable force field to capture the varying

influence of different metal ions in M-MOF-74. Both CO₂ and CH₄ have a similar polarizability¹²⁵ but show a totally different adsorption behavior in the series of M-MOF-74. Previous studies explain the difference with a combination of electrostatic interactions caused by the permanent quadrupole of CO₂ and polarization.⁵¹ As an example, in Figure 9, the predicted adsorption isotherms for CH₄ in the Mg, Co, Ni, and Zn structures are compared to experimental measurements, simulations of Mercado *et al.*,⁶⁴ and the UFF force field.

The data for the remaining structures and the heats of adsorption are provided in Figures S12–S21 (Supporting Information). For low fugacities, simulations with the Mn based structure show an unphysical behavior (compare Figure S17). The DFT-optimized structure does not seem to be perfectly symmetrical. Four adsorption sites close to Mn ions are much stronger for CH₄ than the remaining adsorption sites. These adsorption sites are occupied with CH₄ for all fugacities. To investigate this unusual behavior, we conducted Widom test particle insertions to compute Henry coefficients of CH₄ in Mn-MOF-74 in the limit of infinite dilution condition¹⁴⁹ with and without blocking of these four adsorption sites. In addition, simulations were performed with Lennard-Jones parameters for Mn from the UFF force field and with parameters for Zn from

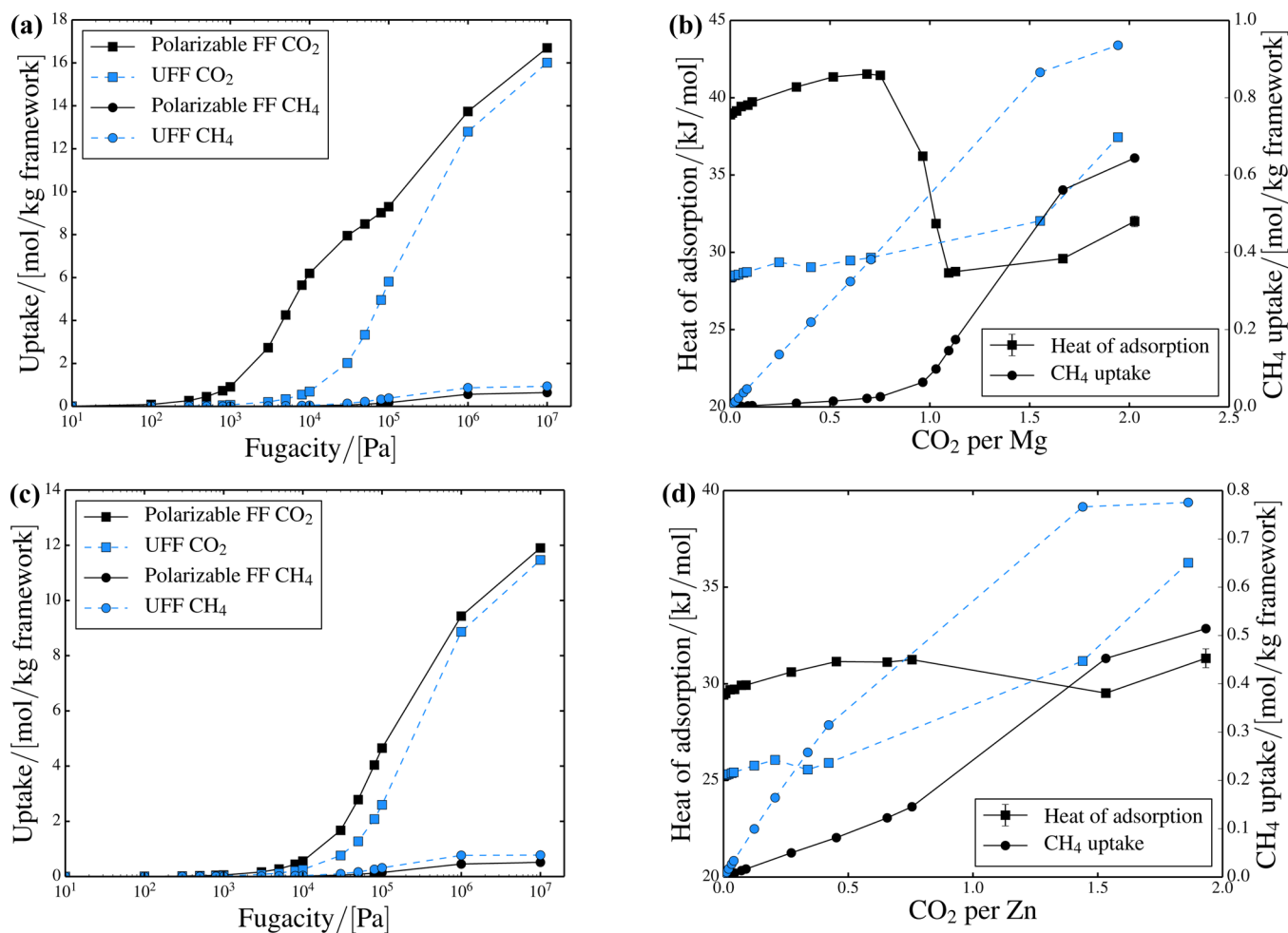


Figure 10. Comparison between the simulation results using the developed polarizable force field (black) and the UFF force field (blue) for an equimolar mixture of CO₂ and CH₄ in the Mg (up) and Zn (down) based structures. (a and c) Adsorption isotherms at 298 K and (b and d) heats of adsorption as a function of uptake in Mg-MOF-74 and Zn-MOF-74, respectively.

the DREIDING force field. The results are presented in the [Supporting Information](#) and show that the four adsorption sites are responsible for the behavior. For the DREIDING parameters, the problem does not occur. The locations of the CH₄ molecules are different due to the different Lennard-Jones parameters and hence interactions are much smaller. Overall, the UFF force field performs better than the developed polarizable force field for CH₄. The polarizable force field underpredicts the uptake for CH₄. This is due to the very simple and crude procedure to determine the force field parameters. Actually, not a single parameter was adjusted for CH₄. CO₂ and CH₄ have a similar polarizability,¹²⁵ but CO₂ is modeled with three interaction sides and CH₄ with only one interaction side. Hence, in our force field, CH₄ has a larger assigned polarizability than CO₂ and the Lennard-Jones energy parameter for CH₄ is more reduced. Besides, CH₄ is modeled without point charges. It is very reassuring that the adsorption isotherms of CH₄ computed with the polarizable force field do not show a distinct inflection which is in agreement with experimental results. As expected by Mishra et al.⁵¹ and observed in this study, the strong interactions in the case of CO₂ are caused by a superposition of static polarity and polarization. Hence, the adsorption of the uncharged CH₄ molecule is far less affected by the different metal ions and a similar behavior is observed for all M-MOF-74 structures. This

is correctly captured by the polarizable force field. In addition, the large CH₄ molecules have a larger distance to the metal ions and hence almost no dipole is induced. It would be straightforward to improve the performance of the polarizable force field by introducing a molecule dependent scaling parameter to account for the unequal number of interaction sides. The simulation results of Mercado et al.⁶⁴ are better than the ones with the UFF force field and the polarizable force field. The experimental adsorption isotherms are well reproduced. However, the same procedure as that for CO₂ was performed to fit the force field parameters for all metal ions separately and the CH₄ uptake is scaled with a factor of 0.85 to account for inaccessible open-metal sites and blocked pores. In comparison to experimental measurements, the UFF, and the polarizable force fields, the curvature of the computed adsorption isotherms for CH₄ of Mercado et al.⁶⁴ seems to be systematically different. The slope of the adsorption isotherm decreases at much lower fugacities. The reason for this behavior should be investigated to further improve the approach. Finally, it is possible to make predictions for the adsorption of mixtures of CH₄ and CO₂ with the developed polarizable force field. Measuring the gas uptake of mixtures in MOFs is more complicated experimentally and is not often done. Therefore, computational prediction of mixtures is very useful to predict the capability of a material to separate gases. In [Figure 10](#), the

predictions of an equimolar mixture of CH₄ and CO₂ are shown for the polarizable force field and compared to predictions based on the UFF force field for Mg-MOF-74 and Zn-MOF-74.

For these structures, the single component adsorption isotherms are reproduced reasonably well and therefore the mixture adsorption isotherms are expected to be predicted reasonably. Predictions of mixtures being adsorbed in the other M-MOF-74 structures can be found in Figures S22–S31 (Supporting Information). Close to the open-metal sites, the interactions between individual guest molecules with the framework are predominant. Inside the channels, the interactions between CO₂ and CH₄ molecules should be dominant which are reproduced well by the TraPPE force field. Due to the strong interactions for CO₂ with the framework, a significantly lower uptake of CH₄ is predicted for Mg-MOF-74 with the polarizable force field. Especially, for low fugacities, mainly CO₂ is adsorbed. Only after most of the Mg sites are saturated CH₄ is taken up. In contrast, the UFF force field predicts an uptake of CH₄ even for very low fugacities. The depicted heats of adsorption show the released heat of adsorption for the mixture as a function of CO₂ uptake per metal ion. For Mg-MOF-74, the initial adsorption behavior is very similar to the one for pure CO₂, because initially almost exclusively CO₂ is adsorbed. In contrast, in Zn-MOF-74, CH₄ is already adsorbed for low fugacities. The trend of the heat of adsorption for the mixture deviates from the ones for the pure components. The Zn based structure seems to be less suitable to separate CO₂ from CH₄. The UFF force field predicts an initially larger uptake of CH₄ in comparison to the polarizable force field. This can partially be attributed to smaller interactions between CH₄ and the framework for Zn-MOF-74 with the polarizable force field.

CONCLUSIONS

The simulations using the developed polarizable force field agree reasonably well with experimental measurements for most of the investigated structures of the M-MOF-74 family. The quality of the predictions for CO₂ is significantly better than with the UFF force field and for most cases comparable to structure specific force fields developed with more elaborated schemes. The polarization energy computed with the polarizable force field shows a behavior similar to the orbital interaction energy determined from DFT calculations. In principle, these energy contributions should be similar if no reaction and no charge transfer takes place. The conducted procedure of first scaling atomic polarizabilities and subsequently adjusting the Lennard-Jones interaction parameters including implicit polarization is simple and requires relatively little effort. The two global scaling factors used here are exclusively tuned for the Mg based structure and CO₂. Hence, the results for the other structures and CH₄ are predictions. For CH₄, no inflection is observed although the largest polarizability is assigned to CH₄. The predictions for CH₄ adsorption could be significantly improved by a molecule specific adjustment of the polarizability. The concept of only considering explicit polarization between guest molecules and the framework and neglecting back-polarization seems to be a well suitable approach to study adsorption phenomena in porous materials. The assumptions considerably enhance the computational performance of Monte Carlo simulations while using polarizable force fields. Actually, the computational time can be similar to Monte Carlo simulations without a polarizable

force field. This is an important assessment, because Monte Carlo simulations are the method of choice for the prediction of adsorption properties in porous materials. Future work will focus on further developing polarizable force fields and deriving a consistent set of parameters from quantum mechanical calculations to avoid fitting to experimental data. We believe that this can lead to force fields with better physical justification and improved transferability. This is crucial for the usage of Monte Carlo simulations for material screening and to make meaningful predictions. Polarizable force fields for Monte Carlo simulations are also promising for other systems with a significant polarization contribution,^{116,150} i.e., water,¹¹⁷ systems including ions,^{116,132} or xylenes.¹¹⁹

ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge on the ACS Publications website at DOI: 10.1021/acs.jpcc.6b12052.

Tables containing force field parameters of the developed polarizable force field and the initial UFF and TraPPE force fields; figures showing separately the simulation results for CO₂, CH₄, and their equimolar mixtures for all considered M-MOF-74 structures (M = Co, Cr, Cu, Fe, Mg, Mn, Ni, Ti, V, Zn); calculated Henry coefficients of CO₂ in the Fe based structure and CH₄ in the Mn based structure in the limit of infinite dilution condition (PDF)

AUTHOR INFORMATION

Corresponding Author

*E-mail: t.j.h.vlugt@tudelft.nl.

ORCID

David Dubbeldam: 0000-0002-4382-1509

Li-Chiang Lin: 0000-0002-2821-9501

Thijs J. H. Vlugt: 0000-0003-3059-8712

Notes

The authors declare no competing financial interest.

ACKNOWLEDGMENTS

This work was sponsored by NWO Exacte Wetenschappen (Physical Sciences) for the use of supercomputer facilities, with financial support from the Nederlandse Organisatie voor Wetenschappelijk Onderzoek (Netherlands Organization for Scientific Research, NWO). T.J.H.V. would like to thank NWO-CW (Chemical Sciences) for a VICI grant.

REFERENCES

- (1) McCarthy, J. J. *Climate Change 2001: Impacts, Adaptation, and Vulnerability: Contribution of Working Group II to the Third Assessment Report of the Intergovernmental Panel on Climate Change*, 1st ed.; Cambridge University Press: Cambridge, U.K., 2001. <http://treconservice.com/onep/wp-content/uploads/2015/01/Impacts-Adaptation-and-Vulnerability.pdf> (accessed Nov 7, 2016).
- (2) Parry, M. L. *Climate Change 2007-Impacts, Adaptation and Vulnerability: Working Group II Contribution to the Fourth Assessment Report of the IPCC*, 1st ed.; Cambridge University Press: New York, 2007; Vol. 4. https://www.ipcc.ch/pdf/assessment-report/ar4/wg2/ar4_wg2_full_report.pdf (accessed Nov 7, 2016).
- (3) Rosenzweig, C.; Karoly, D.; Vicarelli, M.; Neofotis, P.; Wu, Q.; Casassa, G.; Menzel, A.; Root, T. L.; Estrella, N.; Seguin, B. Attributing Physical and Biological Impacts to Anthropogenic Climate Change. *Nature* **2008**, *453*, 353–357.
- (4) Thomas, C. D.; Cameron, A.; Green, R. E.; Bakkenes, M.; Beaumont, L. J.; Collingham, Y. C.; Erasmus, B. F. N.; de Siqueira, M.

- F.; Grainger, A.; Hannah, L.; et al. Extinction Risk from Climate Change. *Nature* **2004**, *427*, 145–148.
- (5) Pacala, S.; Socolow, R. Stabilization Wedges: Solving the Climate Problem for the Next 50 Years with Current Technologies. *Science* **2004**, *305*, 968–972.
- (6) Smit, B.; Reimer, J. R.; Oldenburg, C. M.; Bourg, I. C. *Introduction to Carbon Capture and Sequestration*, 1st ed.; Imperial College Press: London, 2014; Vol. 1.
- (7) Sumida, K.; Rogow, D. L.; Mason, J. A.; McDonald, T. M.; Bloch, E. D.; Herm, Z. R.; Bae, T.-H.; Long, J. R. Carbon Dioxide Capture in Metal-Organic Frameworks. *Chem. Rev.* **2012**, *112*, 724–781.
- (8) Chu, S. Carbon Capture and Sequestration. *Science* **2009**, *325*, 1599.
- (9) Haszeldine, R. S. Carbon Capture and Storage: How Green Can Black Be? *Science* **2009**, *325*, 1647–1652.
- (10) Heinsohn, R. J.; Kabel, R. L. *Sources and Control of Air Pollution: Engineering Principles*, 1st ed.; Prentice Hall: Englewood Cliffs, NJ, 1999.
- (11) Brasseur, G.; Orlando, J.; Tyndall, G., Eds. *Atmospheric Chemistry and Global Change*, 1st ed.; Oxford University Press: New York, 1999.
- (12) Mokhtab, S.; Poe, W. A.; Speight, J. G., Eds. *Handbook of Natural Gas Transmission and Processing*, 1st ed.; Gulf Professional Publishing: Burlington, MA, 2006; pp 261–294.
- (13) Lin, L.-C.; Berger, A. H.; Martin, R. L.; Kim, J.; Swisher, J. A.; Jariwala, K.; Rycroft, C. H.; Bhowan, A. S.; Deem, M. W.; Haranczyk, M. In Silico Screening of Carbon-Capture Materials. *Nat. Mater.* **2012**, *11*, 633–641.
- (14) Huck, J. M.; Lin, L.-C.; Berger, A. H.; Shahrak, M. N.; Martin, R. L.; Bhowan, A. S.; Haranczyk, M.; Reuter, K.; Smit, B. Evaluating Different Classes of Porous Materials for Carbon Capture. *Energy Environ. Sci.* **2014**, *7*, 4132–4146.
- (15) Rowsell, J. L.; Yaghi, O. M. Metal-Organic Frameworks: A New Class of Porous Materials. *Microporous Mesoporous Mater.* **2004**, *73*, 3–14.
- (16) Janiak, C.; Vieth, J. K. MOFs, MILs and More: Concepts, Properties and Applications for Porous Coordination Networks (PCNs). *New J. Chem.* **2010**, *34*, 2366–2388.
- (17) D'Alessandro, D.; Smit, B.; Long, J. Carbon Dioxide Capture: Prospects for New Materials. *Angew. Chem., Int. Ed.* **2010**, *49*, 6058–6082.
- (18) Herm, Z. R.; Swisher, J. A.; Smit, B.; Krishna, R.; Long, J. R. Metal-Organic Frameworks as Adsorbents for Hydrogen Purification and Precombustion Carbon Dioxide Capture. *J. Am. Chem. Soc.* **2011**, *133*, 5664–5667.
- (19) Mason, J. A.; Sumida, K.; Herm, Z. R.; Krishna, R.; Long, J. R. Evaluating Metal-Organic Frameworks for Post-Combustion Carbon Dioxide Capture via Temperature Swing Adsorption. *Energy Environ. Sci.* **2011**, *4*, 3030–3040.
- (20) Bloch, E. D.; Murray, L. J.; Queen, W. L.; Chavan, S.; Maximoff, S. N.; Bigi, J. P.; Krishna, R.; Peterson, V. K.; Grandjean, F.; Long, G. J. Selective Binding of O₂ over N₂ in a Redox-Active Metal-Organic Framework with Open Iron(II) Coordination Sites. *J. Am. Chem. Soc.* **2011**, *133*, 14814–14822.
- (21) Couck, S.; Denayer, J. F. M.; Baron, G. V.; Rémy, T.; Gascon, J.; Kapteijn, F. An Amine-Functionalized MIL-53 Metal-Organic Framework with Large Separation Power for CO₂ and CH₄. *J. Am. Chem. Soc.* **2009**, *131*, 6326–6327.
- (22) McDonald, T. M.; Mason, J. A.; Kong, X.; Bloch, E. D.; Gygi, D.; Dani, A.; Crocella, V.; Giordanino, F.; Odoh, S. O.; Drisdell, W. S.; et al. Cooperative Insertion of CO₂ in Diamine-Appended Metal-Organic Frameworks. *Nature* **2015**, *519*, 303–308.
- (23) Rosi, N. L.; Eckert, J.; Eddaoudi, M.; Vodak, D. T.; Kim, J.; O'Keeffe, M.; Yaghi, O. M. Hydrogen Storage in Microporous Metal-Organic Frameworks. *Science* **2003**, *300*, 1127–1129.
- (24) Zhou, W.; Yildirim, T. Nature and Tunability of Enhanced Hydrogen Binding in Metal-Organic Frameworks with Exposed Transition Metal Sites. *J. Phys. Chem. C* **2008**, *112*, 8132–8135.
- (25) Peng, Y.; Krungleviciute, V.; Eryazici, I.; Hupp, J. T.; Farha, O. K.; Yildirim, T. Methane Storage in Metal-Organic Frameworks: Current Records, Surprise Findings, and Challenges. *J. Am. Chem. Soc.* **2013**, *135*, 11887–11894.
- (26) Collins, D. J.; Zhou, H.-C. Hydrogen Storage in Metal-Organic Frameworks. *J. Mater. Chem.* **2007**, *17*, 3154–3160.
- (27) Liu, H.; Maginn, E. J.; Visser, A. E.; Bridges, N. J.; Fox, E. B. Thermal and Transport Properties of Six Ionic Liquids: An Experimental and Molecular Dynamics Study. *Ind. Eng. Chem. Res.* **2012**, *51*, 7242–7254.
- (28) Wu, H.; Gong, Q.; Olson, D. H.; Li, J. Commensurate Adsorption of Hydrocarbons and Alcohols in Microporous Metal Organic Frameworks. *Chem. Rev.* **2012**, *112*, 836–868.
- (29) Shah, M.; McCarthy, M. C.; Sachdeva, S.; Lee, A. K.; Jeong, H.-K. Current Status of Metal-Organic Framework Membranes for Gas Separations: Promises and Challenges. *Ind. Eng. Chem. Res.* **2012**, *51*, 2179–2199.
- (30) Li, J.-R.; Sculley, J.; Zhou, H.-C. Metal-Organic Frameworks for Separations. *Chem. Rev.* **2012**, *112*, 869–932.
- (31) Czaja, A. U.; Trukhan, N.; Muller, U. Industrial Applications of Metal-Organic Frameworks. *Chem. Soc. Rev.* **2009**, *38*, 1284–1293.
- (32) Lee, J.; Farha, O. K.; Roberts, J.; Scheidt, K. A.; Nguyen, S. T.; Hupp, J. T. Metal-Organic Framework Materials as Catalysts. *Chem. Soc. Rev.* **2009**, *38*, 1450–1459.
- (33) Xiao, D. J.; Bloch, E. D.; Mason, J. A.; Queen, W. L.; Hudson, M. R.; Planas, N.; Borycz, J.; Dzubak, A. L.; Verma, P.; Lee, K.; et al. Oxidation of Ethane to Ethanol by N₂O in a Metal-Organic Framework with Coordinatively Unsaturated Iron(II) Sites. *Nat. Chem.* **2014**, *6*, 590–595.
- (34) Kreno, L. E.; Leong, K.; Farha, O. K.; Allendorf, M.; Dwyne, R. P. V.; Hupp, J. T. Metal-Organic Framework Materials as Chemical Sensors. *Chem. Rev.* **2012**, *112*, 1105–1125.
- (35) Horcajada, P.; Serre, C.; Vallet-Regí, M.; Sebban, M.; Taulelle, F.; Férey, G. Metal-Organic Frameworks as Efficient Materials for Drug Delivery. *Angew. Chem., Int. Ed.* **2006**, *45*, 5974–5978.
- (36) Huxford, R. C.; Rocca, J. D.; Lin, W. Metal-Organic Frameworks as Potential Drug Carriers. *Curr. Opin. Chem. Biol.* **2010**, *14*, 262–268.
- (37) Kitagawa, S.; Kitaura, R.; Noro, S.-i. Functional Porous Coordination Polymers. *Angew. Chem., Int. Ed.* **2004**, *43*, 2334–2375.
- (38) Rao, C. N. R.; Natarajan, S.; Vaidhyanathan, R. Metal Carboxylates with Open Architectures. *Angew. Chem., Int. Ed.* **2004**, *43*, 1466–1496.
- (39) McKinlay, A.; Morris, R.; Horcajada, P.; Férey, G.; Gref, R.; Couvreur, P.; Serre, C. BioMOFs: Metal-Organic Frameworks for Biological and Medical Applications. *Angew. Chem., Int. Ed.* **2010**, *49*, 6260–6266.
- (40) Della Rocca, J.; Liu, D.; Lin, W. Nanoscale Metal-Organic Frameworks for Biomedical Imaging and Drug Delivery. *Acc. Chem. Res.* **2011**, *44*, 957–968.
- (41) Horcajada, P.; Gref, R.; Baati, T.; Allan, P. K.; Maurin, G.; Couvreur, P.; Férey, G.; Morris, R. E.; Serre, C. Metal-Organic Frameworks in Biomedicine. *Chem. Rev.* **2012**, *112*, 1232–1268.
- (42) Park, J.; Kim, H.; Han, S. S.; Jung, Y. Tuning Metal-Organic Frameworks with Open-Metal Sites and Its Origin for Enhancing CO₂ Affinity by Metal Substitution. *J. Phys. Chem. Lett.* **2012**, *3*, 826–829.
- (43) Long, J. R.; Yaghi, O. M. The Pervasive Chemistry of Metal-Organic Frameworks. *Chem. Soc. Rev.* **2009**, *38*, 1213–1214.
- (44) Wilmer, C. E.; Leaf, M.; Lee, C. Y.; Farha, O. K.; Hauser, B. G.; Hupp, J. T.; Snurr, R. Q. Large-Scale Screening of Hypothetical Metal-Organic Frameworks. *Nat. Chem.* **2012**, *4*, 83–89.
- (45) Willems, T. F.; Rycroft, C. H.; Kazi, M.; Meza, J. C.; Haranczyk, M. Algorithms and Tools for High-Throughput Geometry-Based Analysis of Crystalline Porous Materials. *Microporous Mesoporous Mater.* **2012**, *149*, 134–141.
- (46) Queen, W. L.; Hudson, M. R.; Bloch, E. D.; Mason, J. A.; Gonzalez, M. I.; Lee, J. S.; Gygi, D.; Howe, J. D.; Lee, K.; Darwish, T. A.; et al. Comprehensive Study of Carbon Dioxide Adsorption in the Metal-Organic Frameworks M₂(dobdc) (M = Mg, Mn, Fe, Co, Ni, Cu, Zn). *Chem. Sci.* **2014**, *5*, 4569–4581.

- (47) Colon, Y. J.; Snurr, R. Q. High-Throughput Computational Screening of Metal-Organic Frameworks. *Chem. Soc. Rev.* **2014**, *43*, 5735–5749.
- (48) Poloni, R.; Lee, K.; Berger, R. F.; Smit, B.; Neaton, J. B. Understanding Trends in CO₂ Adsorption in Metal-Organic Frameworks with Open-Metal Sites. *J. Phys. Chem. Lett.* **2014**, *5*, 861–865.
- (49) Haldoupis, E.; Borycz, J.; Shi, H.; Vogiatzis, K. D.; Bai, P.; Queen, W. L.; Gagliardi, L.; Siepmann, J. I. Ab Initio Derived Force Fields for Predicting CO₂ Adsorption and Accessibility of Metal Sites in the Metal-Organic Frameworks M-MOF-74 (M = Mn, Co, Ni, Cu). *J. Phys. Chem. C* **2015**, *119*, 16058–16071.
- (50) Caskey, S. R.; Wong-Foy, A. G.; Matzger, A. J. Dramatic Tuning of Carbon Dioxide Uptake via Metal Substitution in a Coordination Polymer with Cylindrical Pores. *J. Am. Chem. Soc.* **2008**, *130*, 10870–10871.
- (51) Mishra, P.; Edubilli, S.; Mandal, B.; Gumma, S. Adsorption Characteristics of Metal-Organic Frameworks Containing Coordinatively Unsaturated Metal Sites: Effect of Metal Cations and Adsorbate Properties. *J. Phys. Chem. C* **2014**, *118*, 6847–6855.
- (52) Yazaydin, A. Ö.; Snurr, R. Q.; Park, T.-H.; Koh, K.; Liu, J.; LeVan, M. D.; Benin, A. I.; Jakubczak, P.; Lanuza, M.; Galloway, D. B. Screening of Metal-Organic Frameworks for Carbon Dioxide Capture from Flue Gas Using a Combined Experimental and Modeling Approach. *J. Am. Chem. Soc.* **2009**, *131*, 18198–18199.
- (53) Poloni, R.; Smit, B.; Neaton, J. B. CO₂ Capture by Metal-Organic Frameworks with van der Waals Density Functionals. *J. Phys. Chem. A* **2012**, *116*, 4957–4964.
- (54) Kim, J.; Martin, R. L.; Rübél, O.; Haranczyk, M.; Smit, B. High-Throughput Characterization of Porous Materials Using Graphics Processing Units. *J. Chem. Theory Comput.* **2012**, *8*, 1684–1693.
- (55) Kim, J.; Lin, L.-C.; Martin, R. L.; Swisher, J. A.; Haranczyk, M.; Smit, B. Large-Scale Computational Screening of Zeolites for Ethane/Ethane Separation. *Langmuir* **2012**, *28*, 11914–11919.
- (56) Kim, J.; Abouelnasr, M.; Lin, L.-C.; Smit, B. Large-Scale Screening of Zeolite Structures for CO₂ Membrane Separations. *J. Am. Chem. Soc.* **2013**, *135*, 7545–7552.
- (57) Watanabe, T.; Sholl, D. S. Accelerating Applications of Metal-Organic Frameworks for Gas Adsorption and Separation by Computational Screening of Materials. *Langmuir* **2012**, *28*, 14114–14128.
- (58) Haldoupis, E.; Nair, S.; Sholl, D. S. Finding MOFs for Highly Selective CO₂/N₂ Adsorption Using Materials Screening Based on Efficient Assignment of Atomic Point Charges. *J. Am. Chem. Soc.* **2012**, *134*, 4313–4323.
- (59) Fang, H.; Demir, H.; Kamakoti, P.; Sholl, D. S. Recent Developments in First-Principles Force Fields for Molecules in Nanoporous Materials. *J. Mater. Chem. A* **2014**, *2*, 274–291.
- (60) Pérez-Pellitero, J.; Amrouche, H.; Siperstein, F.; Pirngruber, G.; Nieto-Draghi, C.; Chaplais, G.; Simon-Masseron, A.; Bazer-Bachi, D.; Peralta, D.; Bats, N. Adsorption of CO₂, CH₄, and N₂ on Zeolitic Imidazolate Frameworks: Experiments and Simulations. *Chem. - Eur. J.* **2010**, *16*, 1560–1571.
- (61) Zang, J.; Nair, S.; Sholl, D. S. Prediction of Water Adsorption in Copper-Based Metal-Organic Frameworks Using Force Fields Derived from Dispersion-Corrected DFT Calculations. *J. Phys. Chem. C* **2013**, *117*, 7519–7525.
- (62) Chen, L.; Grajciar, L.; Nachtigall, P.; Düren, T. Accurate Prediction of Methane Adsorption in a Metal-Organic Framework with Unsaturated Metal Sites by Direct Implementation of an Ab Initio Derived Potential Energy Surface in GCMC Simulation. *J. Phys. Chem. C* **2011**, *115*, 23074–23080.
- (63) Dzubak, A. L.; Lin, L.-C.; Kim, J.; Swisher, J. A.; Poloni, R.; Maximoff, S. N.; Smit, B.; Gagliardi, L. Ab Initio Carbon Capture in Open-Site Metal-Organic Frameworks. *Nat. Chem.* **2012**, *4*, 810–816.
- (64) Mercado, R.; Vlaisavljevich, B.; Lin, L.-C.; Lee, K.; Lee, Y.; Mason, J. A.; Xiao, D. J.; Gonzalez, M. I.; Kapelewski, M. T.; Neaton, J. B.; et al. Force Field Development from Periodic Density Functional Theory Calculations for Gas Separation Applications Using Metal-Organic Frameworks. *J. Phys. Chem. C* **2016**, *120*, 12590–12604.
- (65) Chen, B.; Ockwig, N. W.; Millward, A. R.; Contreras, D. S.; Yaghi, O. M. High H₂ Adsorption in a Microporous Metal-Organic Framework with Open Metal Sites. *Angew. Chem.* **2005**, *117*, 4823–4827.
- (66) Chen, L.; Morrison, C. A.; Düren, T. Improving Predictions of Gas Adsorption in Metal-Organic Frameworks with Coordinatively Unsaturated Metal Sites: Model Potentials, Ab Initio Parameterization, and GCMC Simulations. *J. Phys. Chem. C* **2012**, *116*, 18899–18909.
- (67) Colón, Y. J.; Fairen-Jimenez, D.; Wilmer, C. E.; Snurr, R. Q. High-Throughput Screening of Porous Crystalline Materials for Hydrogen Storage Capacity near Room Temperature. *J. Phys. Chem. C* **2014**, *118*, 5383–5389.
- (68) Campbell, C.; Ferreira-Rangel, C. A.; Fischer, M.; Gomes, J. R. B.; Jorge, M. A. Transferable Model for Adsorption in MOFs with Unsaturated Metal Sites. *J. Phys. Chem. C* **2017**, *121*, 441–458.
- (69) Rappe, A. K.; Casewit, C. J.; Colwell, K. S.; Goddard, W. A.; Skiff, W. M. UFF, a Full Periodic Table Force Field for Molecular Mechanics and Molecular Dynamics Simulations. *J. Am. Chem. Soc.* **1992**, *114*, 10024–10035.
- (70) Potoff, J. J.; Siepmann, J. I. Vapor-Liquid Equilibria of Mixtures Containing Alkanes, Carbon Dioxide, and Nitrogen. *AIChE J.* **2001**, *47*, 1676–1682.
- (71) Borycz, J.; Tiana, D.; Haldoupis, E.; Sung, J. C.; Farha, O. K.; Siepmann, J. I.; Gagliardi, L. CO₂ Adsorption in M-IRMOF-10 (M = Mg, Ca, Fe, Cu, Zn, Ge, Sr, Cd, Sn, Ba). *J. Phys. Chem. C* **2016**, *120*, 12819–12830.
- (72) Duerinck, T.; Bueno-Perez, R.; Vermoortele, F.; Vos, D. E. D.; Calero, S.; Baron, G. V.; Denayer, J. F. M. Understanding Hydrocarbon Adsorption in the UiO-66 Metal-Organic Framework: Separation of (Un)saturated Linear, Branched, Cyclic Adsorbates, Including Stereoisomers. *J. Phys. Chem. C* **2013**, *117*, 12567–12578.
- (73) Bueno-Perez, R.; Martin-Calvo, A.; Gomez-Alvarez, P.; Gutierrez-Sevillano, J. J.; Merkling, P. J.; Vlucht, T. J. H.; van Erp, T. S.; Dubbeldam, D.; Calero, S. Enantioselective Adsorption of Ibuprofen and Lysine in Metal-Organic Frameworks. *Chem. Commun.* **2014**, *50*, 10849–10852.
- (74) Fang, H.; Kamakoti, P.; Zang, J.; Cundy, S.; Paur, C.; Ravikovitch, P. I.; Sholl, D. S. Prediction of CO₂ Adsorption Properties in Zeolites Using Force Fields Derived from Periodic Dispersion-Corrected DFT Calculations. *J. Phys. Chem. C* **2012**, *116*, 10692–10701.
- (75) Lin, L.-C.; Lee, K.; Gagliardi, L.; Neaton, J. B.; Smit, B. Force-Field Development from Electronic Structure Calculations with Periodic Boundary Conditions: Applications to Gaseous Adsorption and Transport in Metal-Organic Frameworks. *J. Chem. Theory Comput.* **2014**, *10*, 1477–1488.
- (76) Borycz, J.; Lin, L.-C.; Bloch, E. D.; Kim, J.; Dzubak, A. L.; Maurice, R.; Semrouni, D.; Lee, K.; Smit, B.; Gagliardi, L. CO₂ Adsorption in Fe₂(dobdc): A Classical Force Field Parameterized from Quantum Mechanical Calculations. *J. Phys. Chem. C* **2014**, *118*, 12230–12240.
- (77) Addicoat, M. A.; Vankova, N.; Akter, I. F.; Heine, T. Extension of the Universal Force Field to Metal-Organic Frameworks. *J. Chem. Theory Comput.* **2014**, *10*, 880–891.
- (78) Vanduyfhuys, L.; Vandenbrande, S.; Verstraelen, T.; Schmid, R.; Waroquier, M.; Van Speybroeck, V. QuickFF: A program for a Quick and Easy derivation of Force Fields for Metal-Organic Frameworks from Ab Initio Input. *J. Comput. Chem.* **2015**, *36*, 1015–1027.
- (79) Wu, H.; Zhou, W.; Yildirim, T. High-Capacity Methane Storage in Metal-Organic Frameworks M₂(dhtp): The Important Role of Open Metal Sites. *J. Am. Chem. Soc.* **2009**, *131*, 4995–5000.
- (80) Lee, K.; Howe, J. D.; Lin, L.-C.; Smit, B.; Neaton, J. B. Small-Molecule Adsorption in Open-Site Metal-Organic Frameworks: A Systematic Density Functional Theory Study for Rational Design. *Chem. Mater.* **2015**, *27*, 668–678.
- (81) Dietzel, P. D.; Besikiotis, V.; Blom, R. Application of Metal-Organic Frameworks with Coordinatively Unsaturated Metal Sites in Storage and Separation of Methane and Carbon Dioxide. *J. Mater. Chem.* **2009**, *19*, 7362–7370.

- (82) Dincă, M.; Long, J. Hydrogen Storage in Microporous Metal-Organic Frameworks with Exposed Metal Sites. *Angew. Chem., Int. Ed.* **2008**, *47*, 6766–6779.
- (83) Britt, D.; Furukawa, H.; Wang, B.; Glover, T. G.; Yaghi, O. M. Highly Efficient Separation of Carbon Dioxide by a Metal-Organic Framework Replete with Open Metal Sites. *Proc. Natl. Acad. Sci. U. S. A.* **2009**, *106*, 20637–20640.
- (84) Simmons, J. M.; Wu, H.; Zhou, W.; Yildirim, T. Carbon Capture in Metal-Organic Frameworks-A Comparative Study. *Energy Environ. Sci.* **2011**, *4*, 2177–2185.
- (85) Krishna, R.; van Baten, J. M. Investigating the Potential of MgMOF-74 Membranes for CO₂ Capture. *J. Membr. Sci.* **2011**, *377*, 249–260.
- (86) Chavan, S.; Vitillo, J. G.; Groppo, E.; Bonino, F.; Lamberti, C.; Dietzel, P. D. C.; Bordiga, S. CO Adsorption on CPO-27-Ni Coordination Polymer: Spectroscopic Features and Interaction Energy. *J. Phys. Chem. C* **2009**, *113*, 3292–3299.
- (87) Bloch, E. D.; Queen, W. L.; Krishna, R.; Zadrozny, J. M.; Brown, C. M.; Long, J. R. Hydrocarbon Separations in a Metal-Organic Framework with Open Iron(II) Coordination Sites. *Science* **2012**, *335*, 1606–1610.
- (88) He, Y.; Krishna, R.; Chen, B. Metal-Organic Frameworks with Potential for Energy-Efficient Adsorptive Separation of Light Hydrocarbons. *Energy Environ. Sci.* **2012**, *5*, 9107–9120.
- (89) Bao, Z.; Alnemrat, S.; Yu, L.; Vasiliev, I.; Ren, Q.; Lu, X.; Deng, S. Adsorption of Ethane, Ethylene, Propane, and Propylene on a Magnesium-Based Metal-Organic Framework. *Langmuir* **2011**, *27*, 13554–13562.
- (90) Geier, S. J.; Mason, J. A.; Bloch, E. D.; Queen, W. L.; Hudson, M. R.; Brown, C. M.; Long, J. R. Selective Adsorption of Ethylene over Ethane and Propylene over Propane in the Metal-Organic Frameworks M₂(dobdc) (M = Mg, Mn, Fe, Co, Ni, Zn). *Chem. Sci.* **2013**, *4*, 2054–2061.
- (91) Wu, H.; Simmons, J. M.; Srinivas, G.; Zhou, W.; Yildirim, T. Adsorption Sites and Binding Nature of CO₂ in Prototypical Metal-Organic Frameworks: A Combined Neutron Diffraction and First-Principles Study. *J. Phys. Chem. Lett.* **2010**, *1*, 1946–1951.
- (92) Queen, W. L.; Brown, C. M.; Britt, D. K.; Zajdel, P.; Hudson, M. R.; Yaghi, O. M. Site-Specific CO₂ Adsorption and Zero Thermal Expansion in an Anisotropic Pore Network. *J. Phys. Chem. C* **2011**, *115*, 24915–24919.
- (93) Drisdell, W. S.; Poloni, R.; McDonald, T. M.; Long, J. R.; Smit, B.; Neaton, J. B.; Prendergast, D.; Kortright, J. B. Probing Adsorption Interactions in Metal-Organic Frameworks Using X-ray Spectroscopy. *J. Am. Chem. Soc.* **2013**, *135*, 18183–18190.
- (94) Tan, K.; Zuluaga, S.; Gong, Q.; Gao, Y.; Nijem, N.; Li, J.; Thonhauser, T.; Chabal, Y. J. Competitive Coadsorption of CO₂ with H₂O, NH₃, SO₂, NO, NO₂, N₂, O₂, and CH₄ in M-MOF-74 (M = Mg, Co, Ni): The Role of Hydrogen Bonding. *Chem. Mater.* **2015**, *27*, 2203–2217.
- (95) Kong, X.; Scott, E.; Ding, W.; Mason, J. A.; Long, J. R.; Reimer, J. A. CO₂ Dynamics in a Metal-Organic Framework with Open Metal Sites. *J. Am. Chem. Soc.* **2012**, *134*, 14341–14344.
- (96) Lin, L.-C.; Kim, J.; Kong, X.; Scott, E.; McDonald, T. M.; Long, J. R.; Reimer, J. A.; Smit, B. Understanding CO₂ Dynamics in Metal-Organic Frameworks with Open Metal Sites. *Angew. Chem.* **2013**, *125*, 4506–4509.
- (97) Gutiérrez-Sevillano, J. J.; Dubbeldam, D.; Rey, F.; Valencia, S.; Palomino, M.; Martín-Calvo, A.; Calero, S. Analysis of the ITQ-12 Zeolite Performance in Propane-Propylene Separations Using a Combination of Experiments and Molecular Simulations. *J. Phys. Chem. C* **2010**, *114*, 14907–14914.
- (98) Dubbeldam, D.; Walton, K. S. On the Application of Classical Molecular Simulations of Adsorption in Metal-Organic Frameworks. In *Metal-Organic Frameworks: Materials Modeling towards Engineering Applications*, 1st ed.; CRC Press: Boca Raton, FL, 2015; pp 53–112.
- (99) Pham, T.; Forrest, K. A.; Hogan, A.; McLaughlin, K.; Belof, J. L.; Eckert, J.; Space, B. Simulations of Hydrogen Sorption in rht-MOF-1: Identifying the Binding Sites Through Explicit Polarization and Quantum Rotation Calculations. *J. Mater. Chem. A* **2014**, *2*, 2088–2100.
- (100) Cirera, J.; Sung, J. C.; Howland, P. B.; Paesani, F. The Effects of Electronic Polarization on Water Adsorption in Metal-Organic Frameworks: H₂O in MIL-53(Cr). *J. Chem. Phys.* **2012**, *137*, 054704.
- (101) Wiers, B. M.; Foo, M.-L.; Balsara, N. P.; Long, J. R. A Solid Lithium Electrolyte via Addition of Lithium Isopropoxide to a Metal-Organic Framework with Open Metal Sites. *J. Am. Chem. Soc.* **2011**, *133*, 14522–14525.
- (102) Hwang, Y.; Hong, D.-Y.; Chang, J.-S.; Jung, S.; Seo, Y.-K.; Kim, J.; Vimont, A.; Daturi, M.; Serre, C.; Férey, G. Amine Grafting on Coordinatively Unsaturated Metal Centers of MOFs: Consequences for Catalysis and Metal Encapsulation. *Angew. Chem., Int. Ed.* **2008**, *47*, 4144–4148.
- (103) Demessence, A.; D'Alessandro, D. M.; Foo, M. L.; Long, J. R. Strong CO₂ Binding in a Water-Stable, Triazolate-Bridged Metal-Organic Framework Functionalized with Ethylenediamine. *J. Am. Chem. Soc.* **2009**, *131*, 8784–8786.
- (104) Yu, D.; Yazaydin, A. Ö.; Lane, J. R.; Dietzel, P. D. C.; Snurr, R. Q. A Combined Experimental and Quantum Chemical Study of CO₂ Adsorption in the Metal-Organic Framework CPO-27 with Different Metals. *Chem. Sci.* **2013**, *4*, 3544–3556.
- (105) Schmidt, J. R.; Yu, K.; McDaniel, J. G. Transferable Next-Generation Force Fields from Simple Liquids to Complex. *Acc. Chem. Res.* **2015**, *48*, 548–556.
- (106) Odoh, S. O.; Cramer, C. J.; Truhlar, D. G.; Gagliardi, L. Quantum-Chemical Characterization of the Properties and Reactivities of Metal-Organic Frameworks. *Chem. Rev.* **2015**, *115*, 6051–6111.
- (107) Mayo, S. L.; Olafson, B. D.; Goddard, W. A. DREIDING: A Generic Force Field for Molecular Simulations. *J. Phys. Chem.* **1990**, *94*, 8897–8909.
- (108) Jorgensen, W. L.; Madura, J. D.; Swenson, C. J. Optimized Intermolecular Potential Functions for Liquid Hydrocarbons. *J. Am. Chem. Soc.* **1984**, *106*, 6638–6646.
- (109) Dubbeldam, D.; Calero, S.; Vlugt, T. J. H.; Krishna, R.; Maesen, T. L. M.; Beerdse, E.; Smit, B. Force Field Parametrization through Fitting on Inflection Points in Isotherms. *Phys. Rev. Lett.* **2004**, *93*, 088302.
- (110) Halgren, T. A.; Damm, W. Polarizable Force Fields. *Curr. Opin. Struct. Biol.* **2001**, *11*, 236–242.
- (111) Yu, H.; van Gunsteren, W. F. Accounting for Polarization in Molecular Simulation. *Comput. Phys. Commun.* **2005**, *172*, 69–85.
- (112) Warshel, A.; Kato, M.; Pislakov, A. V. Polarizable Force Fields: History, Test Cases, and Prospects. *J. Chem. Theory Comput.* **2007**, *3*, 2034–2045.
- (113) Cieplak, P.; Dupradeau, F.-Y.; Duan, Y.; Wang, J. Polarization Effects in Molecular Mechanical Force Fields. *J. Phys.: Condens. Matter* **2009**, *21*, 333102.
- (114) Lopes, P. E. M.; Roux, B.; MacKerell, A. D. Molecular Modeling and Dynamics Studies with Explicit Inclusion of Electronic Polarizability: Theory and Applications. *Theor. Chem. Acc.* **2009**, *124*, 11–28.
- (115) Chen, B.; Potoff, J. J.; Siepmann, J. I. Adiabatic Nuclear and Electronic Sampling Monte Carlo Simulations in the Gibbs Ensemble: Application to Polarizable Force Fields for Water. *J. Phys. Chem. B* **2000**, *104*, 2378–2390.
- (116) Antila, H. S.; Salonen, E. Polarizable Force Fields. In *Biomolecular Simulations*, 1st ed.; Humana Press: Clifton, NJ, 2013; pp 215–241.
- (117) Rick, S. W.; Stuart, S. J. Potentials and Algorithms for Incorporating Polarizability in Computer Simulations. *Reviews in Computational Chemistry*, 1st ed.; John Wiley & Sons, Inc.: Hoboken, NJ, 2003; Vol. 18, pp 89–146.
- (118) Becker, T. M.; Dubbeldam, D.; Lin, L.-C.; Vlugt, T. J. H. Investigating Polarization Effects of CO₂ Adsorption in MgMOF-74. *J. Comput. Sci.* **2016**, *15*, 86–94.
- (119) Lachet, V.; Boutin, A.; Tavittian, B.; Fuchs, A. Computational Study of p-Xylene/m-Xylene Mixtures Adsorbed in NaY Zeolite. *J. Phys. Chem. B* **1998**, *102*, 9224–9233.

- (120) Keskin, S.; Liu, J.; Rankin, R. B.; Johnson, J. K.; Sholl, D. S. Progress, Opportunities, and Challenges for Applying Atomically Detailed Modeling to Molecular Adsorption and Transport in Metal-Organic Framework Materials. *Ind. Eng. Chem. Res.* **2009**, *48*, 2355–2371.
- (121) Goj, A.; Sholl, D. S.; Akten, E. D.; Kohen, D. Atomistic Simulations of CO₂ and N₂ Adsorption in Silica Zeolites: The Impact of Pore Size and Shape. *J. Phys. Chem. B* **2002**, *106*, 8367–8375.
- (122) Gee, J. A.; Chung, J.; Nair, S.; Sholl, D. S. Adsorption and Diffusion of Small Alcohols in Zeolitic Imidazolate Frameworks ZIF-8 and ZIF-90. *J. Phys. Chem. C* **2013**, *117*, 3169–3176.
- (123) Zeng, Y.; Moghadam, P. Z.; Snurr, R. Q. Pore Size Dependence of Adsorption and Separation of Thiophene/Benzene Mixtures in Zeolites. *J. Phys. Chem. C* **2015**, *119*, 15263–15273.
- (124) Shannon, R. D. Dielectric Polarizabilities of Ions in Oxides and Fluorides. *J. Appl. Phys.* **1993**, *73*, 348–366.
- (125) van Duijnen, P. T.; Swart, M. Molecular and Atomic Polarizabilities: Thole's Model Revisited. *J. Phys. Chem. A* **1998**, *102*, 2399–2407.
- (126) Pellenq, R. J.-M.; Nicholson, D. Intermolecular Potential Function for the Physical Adsorption of Rare Gases in Silicalite. *J. Phys. Chem.* **1994**, *98*, 13339–13349.
- (127) Tu, Y.; Laaksonen, A. The Electronic Properties of Water Molecules in Water Clusters and Liquid Water. *Chem. Phys. Lett.* **2000**, *329*, 283–288.
- (128) in het in het Panhuis, M.; Popelier, P. L. A.; Munn, R. W.; Ángyán, J. G. Distributed Polarizability of the Water Dimer: Field-Induced Charge Transfer along the Hydrogen Bond. *J. Chem. Phys.* **2001**, *114*, 7951–7961.
- (129) Bonin, K. D.; Kresin, V. V. *Electric-Dipole Polarizabilities of Atoms, Molecules, and Clusters*, 1st ed.; World Scientific: Singapore, 1997.
- (130) Baker, C.; MacKerell, J.; Alexander, D. Polarizability Rescaling and Atom-Based Thole Scaling in the CHARMM Drude Polarizable Force Field for Ethers. *J. Mol. Model.* **2010**, *16*, 567–576.
- (131) Vosmeer, C. R.; Rustenburg, A. S.; Rice, J. E.; Horn, H. W.; Swope, W. C.; Geerke, D. P. QM/MM-Based Fitting of Atomic Polarizabilities for Use in Condensed-Phase Biomolecular Simulation. *J. Chem. Theory Comput.* **2012**, *8*, 3839–3853.
- (132) Cieplak, P.; Kollman, P. Monte Carlo Simulation of Aqueous Solutions of Li⁺ and Na⁺ Using Many-Body Potentials. Coordination Numbers, Ion Solvation Enthalpies, and the Relative Free Energy of Solvation. *J. Chem. Phys.* **1990**, *92*, 6761–6767.
- (133) Dubbeldam, D.; Calero, S.; Ellis, D. E.; Snurr, R. Q. RASPA: Molecular Simulation Software for Adsorption and Diffusion in Flexible Nanoporous Materials. *Mol. Simul.* **2016**, *42*, 81–101.
- (134) Dubbeldam, D.; Torres-Knoop, A.; Walton, K. S. On the Inner Workings of Monte Carlo Codes. *Mol. Simul.* **2013**, *39*, 1253–1292.
- (135) Schnabel, T.; Vrabec, J.; Hasse, H. Unlike Lennard-Jones Parameters for Vapor-Liquid Equilibria. *J. Mol. Liq.* **2007**, *135*, 170–178.
- (136) Frenkel, D.; Smit, B. *Understanding Molecular Simulation*, 2nd ed.; Academic Press: San Diego, CA, 2002.
- (137) Peng, D.-Y.; Robinson, D. B. A New Two-Constant Equation of State. *Ind. Eng. Chem. Fundam.* **1976**, *15*, 59–64.
- (138) te Velde, G.; Bickelhaupt, F. M.; Baerends, E. J.; Fonseca Guerra, C.; van Gisbergen, S. J. A.; Snijders, J. G.; Ziegler, T. Chemistry with ADF. *J. Comput. Chem.* **2001**, *22*, 931–967.
- (139) Baerends, E. J.; Ziegler, T.; Atkins, A. J.; Autschbach, J.; Bashford, D.; Bérces, A.; Bickelhaupt, F. M.; Bo, C.; Boerrigter, P. M.; Cavallo, L.; et al. ADF 2013, SCM, Theoretical Chemistry, Vrije Universiteit, Amsterdam, The Netherlands, <http://www.scm.com>.
- (140) Becke, A. D. Density-Functional Thermochemistry. III. The Role of Exact Exchange. *J. Chem. Phys.* **1993**, *98*, 5648–5652.
- (141) Lee, C.; Yang, W.; Parr, R. G. Development of the Colle-Salvetti Correlation-Energy Formula Into a Functional of the Electron Density. *Phys. Rev. B: Condens. Matter Mater. Phys.* **1988**, *37*, 785–789.
- (142) Vosko, S. H.; Wilk, L.; Nusair, M. Accurate Spin-Dependent Electron Liquid Correlation Energies for Local Spin Density Calculations: A Critical Analysis. *Can. J. Phys.* **1980**, *58*, 1200–1211.
- (143) Stephens, P. J.; Devlin, F. J.; Chabalowski, C. F.; Frisch, M. J. Ab Initio Calculation of Vibrational Absorption and Circular Dichroism Spectra Using Density Functional Force Fields. *J. Phys. Chem.* **1994**, *98*, 11623–11627.
- (144) Grimme, S.; Antony, J.; Ehrlich, S.; Krieg, H. A Consistent and Accurate Ab Initio Parametrization of Density Functional Dispersion Correction (DFT-D) for the 94 Elements H-Pu. *J. Chem. Phys.* **2010**, *132*, 154104.
- (145) Ziegler, T.; Rauk, A. Carbon Monoxide, Carbon Monosulfide, Molecular Nitrogen, Phosphorus Trifluoride, and Methyl Isocyanide as Sigma Donors and Pi Acceptors. A Theoretical Study by the Hartree-Fock-Slater Transition-State Method. *Inorg. Chem.* **1979**, *18*, 1755–1759.
- (146) Ziegler, T.; Rauk, A. A Theoretical Study of the Ethylene-Metal Bond in Complexes between Copper(1+), Silver(1+), Gold(1+), Platinum(0) or Platinum(2+) and Ethylene, Based on the Hartree-Fock-Slater Transition-State Method. *Inorg. Chem.* **1979**, *18*, 1558–1565.
- (147) Bickelhaupt, F. M.; Baerends, E. J. Kohn-Sham Density Functional Theory: Predicting and Understanding Chemistry. *Reviews in Computational Chemistry*, 1st ed.; John Wiley & Sons, Inc.: New York, 2000; Vol. 15, pp 1–86.
- (148) Lee, K.; Isley, W. C.; Dzubak, A. L.; Verma, P.; Stoneburner, S. J.; Lin, L.-C.; Howe, J. D.; Bloch, E. D.; Reed, D. A.; Hudson, M. R.; et al. Design of a Metal-Organic Framework with Enhanced Back Bonding for Separation of N₂ and CH₄. *J. Am. Chem. Soc.* **2014**, *136*, 698–704.
- (149) Vlugt, T. J. H.; García-Pérez, E.; Dubbeldam, D.; Ban, S.; Calero, S. Computing the Heat of Adsorption Using Molecular Simulations: The Effect of Strong Coulombic Interactions. *J. Chem. Theory Comput.* **2008**, *4*, 1107–1118.
- (150) Cieplak, P.; Caldwell, J.; Kollman, P. Molecular Mechanical Models for Organic and Biological Systems going Beyond the Atom Centered Two Body Additive Approximation: Aqueous Solution Free Energies of Methanol and n-Methyl Acetamide, Nucleic Acid Base, and Amide Hydrogen Bonding and Chloroform/Water Partition Coefficients of the Nucleic Acid Bases. *J. Comput. Chem.* **2001**, *22*, 1048–1057.