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# Gas Phase Sensing of Alcohols by Metal Organic Framework–Polymer Composite Materials

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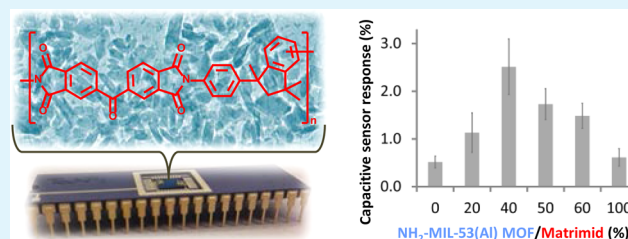
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## Supporting Information

**ABSTRACT:** Affinity layers play a crucial role in chemical sensors for the selective and sensitive detection of analytes. Here, we report the use of composite affinity layers containing Metal Organic Frameworks (MOFs) in a polymeric matrix for sensing purposes. Nanoparticles of NH<sub>2</sub>-MIL-53(Al) were dispersed in a Matrimid polymer matrix with different weight ratios (0–100 wt %) and drop-casted on planar capacitive transducer devices. These coated devices were electrically analyzed using impedance spectroscopy and investigated for their sensing properties toward the detection of a series of alcohols and water in the gas phase. The measurements indicated a reversible and reproducible response in all devices. Sensor devices containing 40 wt % NH<sub>2</sub>-MIL-53(Al) in Matrimid showed a maximum response for methanol and water. The sensor response time slowed down with increasing MOF concentration until 40 wt %. The half time of saturation response ( $\tau_{0.5}$ ) increased by  $\sim 1.75$  times for the 40 wt % composition compared to devices coated with Matrimid only. This is attributed to polymer rigidification near the MOF/polymer interface. Higher MOF loadings ( $\geq 50$  wt %) resulted in brittle coatings with a response similar to the 100 wt % MOF coating. Cross-sensitivity studies showed the ability to kinetically distinguish between the different alcohols with a faster response for methanol and water compared to ethanol and 2-propanol. The observed higher affinity of the pure Matrimid polymer toward methanol compared to water allows also for a higher uptake of methanol in the composite matrices. Also, as indicated by the sensing studies with a mixture of water and methanol, the methanol uptake is independent of the presence of water up to 6000 ppm of water. The NH<sub>2</sub>-MIL-53(Al) MOFs dispersed in the Matrimid matrix show a sensitive and reversible capacitive response, even in the presence of water. By tuning the precise compositions, the affinity kinetics and overall affinity can be tuned, showing the promise of this type of chemical sensors.

**KEYWORDS:** metal organic frameworks, composites, mixed matrix membranes, impedance spectroscopy, capacitive detection, gas sensors



## 1. INTRODUCTION

Detection of analytes in the gas phase is important in environmental, security, health, food, and industrial processes requiring detection and quantification in a fast, accurate, and reversible manner.<sup>1–3</sup> In recent years, studies have focused on developing chemical coatings as a selective affinity layer in sensors using nanomaterials<sup>4,5</sup> and polymeric films.<sup>6,7</sup> The uptake of analytes in the affinity layer leads to changes in the physical properties of the films which can be monitored by the transducer beneath. Often, metal oxides are used as the affinity layer which usually require high operation temperatures ( $>200$  °C) and do show a significant cross-sensitivity in most cases.<sup>8,9</sup>

Metal organic frameworks (MOFs) are a new emerging class of materials of which their specific affinity properties can be attractive for use as selective substrates in chemical gas phase sensors operating at low temperature.<sup>2,10,11</sup> These materials

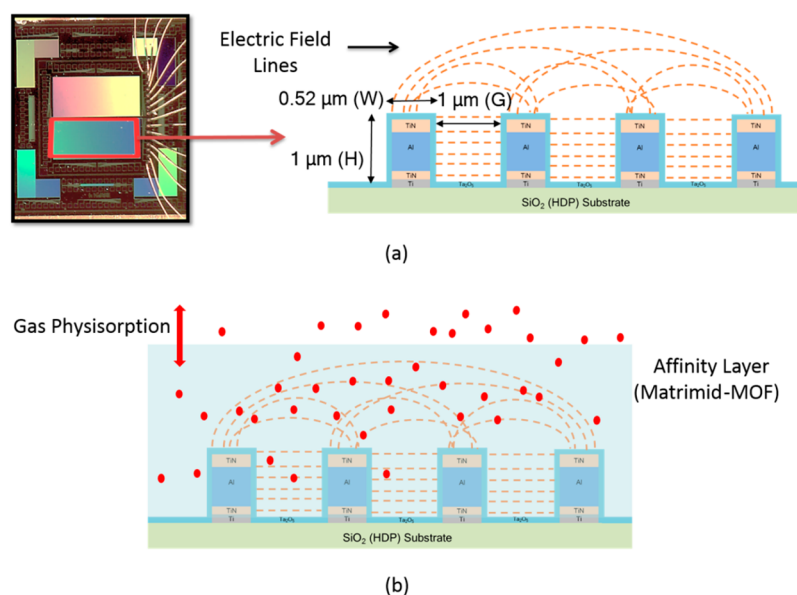
consist of metal ions linked by organic ligand molecules forming porous, crystalline frameworks.<sup>11,12</sup> Their high porosity and tunability of their structure and selectivity make them promising for applications not only for sensing, but also for gas separation, storage, and catalysis.<sup>11–14</sup>

To date, sensor studies using MOFs have focused mainly on changing their luminescence properties.<sup>15,16</sup> Recent advances in the preparation of thin films of MOFs have opened their potential in the sensor industry beyond luminescence-based sensors.<sup>13,17</sup> These recent studies are mainly based on techniques like Quartz-crystal Microbalance (QCM), Surface Acoustic Waves (SAW), and cantilevers, which utilize changes

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**Figure 1.** (a) The sensor device consisting of meander-patterned electrodes with height ( $H$ ), width ( $W$ ) and gap ( $G$ ) of  $1\ \mu\text{m}$ ,  $0.52\ \mu\text{m}$  and  $1\ \mu\text{m}$ , respectively. (b) Schematic representation of the device with the higher dielectric affinity layer and gas physisorption within this layer.

in mass, mechanical properties for analyte detection in MOF-based sensor devices.<sup>17</sup> Also, studies utilizing changes in the electrical properties of MOFs have shown promising application of MOFs as selective affinity layers for sensors.<sup>18–20</sup>

Specifically, MOF-based capacitive sensors offer advantages over resistive sensors due to the insulating nature of MOFs. Compared to, e.g., field-effect transistors, capacitive sensors can be fabricated relatively easily at low cost on a large variety of substrates. Furthermore, they have a low energy consumption and allow different electrode architectures to reduce noise and improve sensitivity.

Additionally, the use of polymeric composite materials with MOFs can combine the properties of both polymers and MOFs for obtaining superior sensing performance. This has also been exploited in gas separation applications.<sup>21,22</sup> The enhanced compatibility of polymers and MOFs contribute also to their promising role in producing new sensor selector layers.<sup>21–23</sup> Studies with polymeric sensors containing filler molecules such as metal oxides, carbon nanotubes, and graphene have already shown improvement of the intrinsic sensing properties of polymers.<sup>24–26</sup> Moreover, the use of composite materials can allow for easy and cost-effective integration with semiconductor technologies utilizing existing principles for synthesizing polymeric coatings.<sup>27,28</sup>

Recently, we reported a simple methodology for developing polymer-MOF composite films using Matrimid 5218 and  $\text{NH}_2\text{-MIL-53(Al)}$  for sensing applications.<sup>29</sup> Matrimid 5218 belongs to the class of fully imidized polyimides with a high thermal and mechanical stability<sup>30,31</sup> and  $\text{NH}_2\text{-MIL-53(Al)}$  is one of the most well-studied MOFs with aluminum as the metal ion and amino-terephthalic acid as the organic ligand.<sup>31,32</sup> It was observed that incorporation of these MOF particles inside the Matrimid polymer matrix enhances the sensitivity of the sensor devices. Matrimid and  $\text{NH}_2\text{-MIL-53(Al)}$  have also shown significant compatibility during membrane separation studies leading to superior gas separation performances.<sup>31</sup> The presence of these MOF particles inside the polymeric matrix alter their physical and chemical properties.<sup>29,31</sup> For example, it has been observed from membrane studies that variations of the

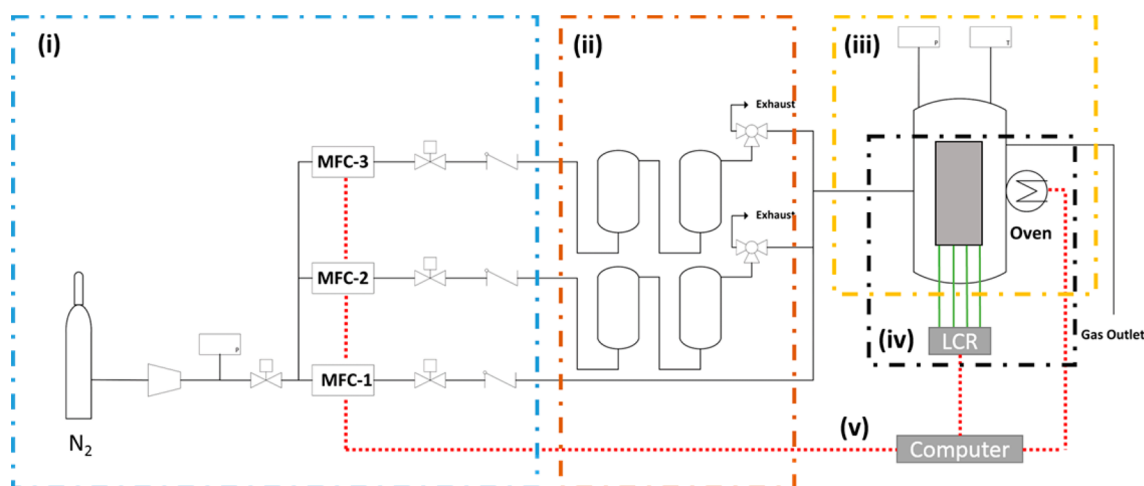
concentration of MOF particles can lead to alterations in the gas transport and adsorption properties.<sup>31,33</sup> Furthermore, upon blending the overall dielectric constant—and hence the electrical polarizability—changes.<sup>34</sup> These differences in the properties with varying MOF concentration can be quite significant for a pair of MOF and polymer.

In this study, we focused on developing planar transducer sensor devices by the deposition of composite films consisting of different amounts of  $\text{NH}_2\text{-MIL-53(Al)}$  in Matrimid. These devices are meander-patterned planar aluminum electrodes having a width ( $W$ ) of  $0.52\ \mu\text{m}$ , a gap ( $G$ ) of  $1\ \mu\text{m}$ , and an electrode area of  $\sim 2.1\ \text{mm}^2$  (Figure 1). About 95% of the electrical field lines stay within the distance equal to the sum of  $W$  and  $G$  in a perpendicular orientation.<sup>35–37</sup> Our bare electrodes show a capacitance of  $\sim 22\ \text{pF}$ . After deposition of polymer-MOF affinity layers of  $10\ \mu\text{m}$  in thickness onto our electrode, the capacitance increases to  $28\text{--}36\ \text{pF}$  as a result of the higher dielectric constant of the deposited layer ( $\epsilon_r = \sim 3.5\text{--}4.5$ ) as compared to air ( $\epsilon_r = 1$ ), as previously discussed.<sup>29</sup> If analyte molecules are adsorbed in the polymer-MOF matrix, then the overall dielectric constant will also change, resulting in a capacitance change.<sup>29</sup> We have performed such sensing experiments using methanol and water as the analyte, varying the polymer-MOF composition ratio. In addition, cross-sensitivity studies with ethanol and 2-propanol and temperature-dependent response studies were performed.

## 2. EXPERIMENTAL SECTION

**2.1. Materials.** The polyimide (PI) Matrimid 5218 was obtained from Huntsman Advanced Materials and was used as the polymeric matrix. Methanol, ethanol, and 2-propanol were obtained from Sigma-Aldrich at highest purity.  $\text{N}_2$  gas (99.99%) was obtained from Linde Gas and used as carrier gas in our sensor experiments. All other chemicals were obtained from Sigma-Aldrich and used as received.

**2.2. Device Fabrication.** The fabrication of the devices was performed as described previously.<sup>29</sup> Briefly, the active electrode area consisted of aluminum metal stacked between TiN layers over underlying dielectric layer of  $\text{SiO}_2$  to mimic the final layer of a typical  $140\ \text{nm}$  CMOS platform. Devices were processed to obtain meander-patterned planar aluminum electrodes by photolithography. These



**Figure 2.** Schematic of the measurement setup consisting of a (i) gas mixing unit with mass-flow controllers, (ii) vapor generation system with series of bubblers, (iii) temperature-controlled measurement chamber, (iv) electrical measurement unit, and (v) instrument control and data logging using LabVIEW and computer.

electrodes consist of 4 parallel electrodes with a meander pattern with line width ( $W$ ) of  $0.52\ \mu\text{m}$ , a gap ( $G$ ) of  $1\ \mu\text{m}$  and a total electrode area of  $\sim 2.1\ \text{mm}^2$ . First,  $\text{Si}_3\text{N}_4$  was deposited on the backside of the p-doped Si wafer to avoid contamination of the Si during the wet cleaning steps. A high-density plasma CVD silicon dioxide (with a thickness of  $3\ \mu\text{m}$ ) was deposited to represent the intermetal dielectric, followed by the deposition of the metal stack. This metal stack consists of a  $25/25\ \text{nm}$  Ti/TiN adhesion bilayer at the bottom and followed by a  $800\ \text{nm}$  thick layer of aluminum and a  $50\ \text{nm}$  layer of TiN at the top. The Ti/TiN and TiN layers were introduced to improve the adhesion, to decrease reflectivity, and also to function as a stopping layer during the etching processes. Afterward, the Al metal layer for the device was patterned by lithography targeting the required width and spacing dimensions. After photoresist patterning, the Al lines were dry etched to obtain 4 parallel meander-patterned aluminum lines. Subsequently, a  $\text{Ta}_2\text{O}_5$  line of  $40\ \text{nm}$  was deposited by CVD to serve as a moisture diffusion barrier.

**2.3. Synthesis of the MOF Nanoparticles.**  $\text{NH}_2\text{-MIL-53(Al)}$  nanoparticles were synthesized by hydrothermal treatment using 2-aminoterephthalic acid as organic ligand and aluminum nitrate using the procedure previously described.<sup>22,38</sup> Therefore,  $1.902\ \text{g}$  ( $= 10.5\ \text{mmol}$ ) 2-aminoterephthalic acid was dissolved in  $10.5\ \text{mL}$   $2\ \text{M}$  aqueous NaOH solution after which the volume was increased to  $75\ \text{mL}$  using distilled water.  $3.935\ \text{g}$   $\text{Al}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$  ( $= 10.5\ \text{mmol}$ ) was dissolved in  $75\ \text{mL}$  distilled water in a separate tube. Both solutions were mixed and stirred under reflux conditions ( $393\ \text{K}$ ) for 3 days. The formed MOF particles were filtered under vacuum and washed with 96% ethanol twice. The material was then washed in DMF ( $\sim 100\ \text{mL}$ ) and methanol ( $\sim 100\ \text{mL}$ ) consecutively at  $403$  and  $363\ \text{K}$  overnight, respectively to efficiently remove any unreacted ligand. Finally, the powder was filtered and washed with ethanol twice and dried at  $373\ \text{K}$  in vacuum overnight.

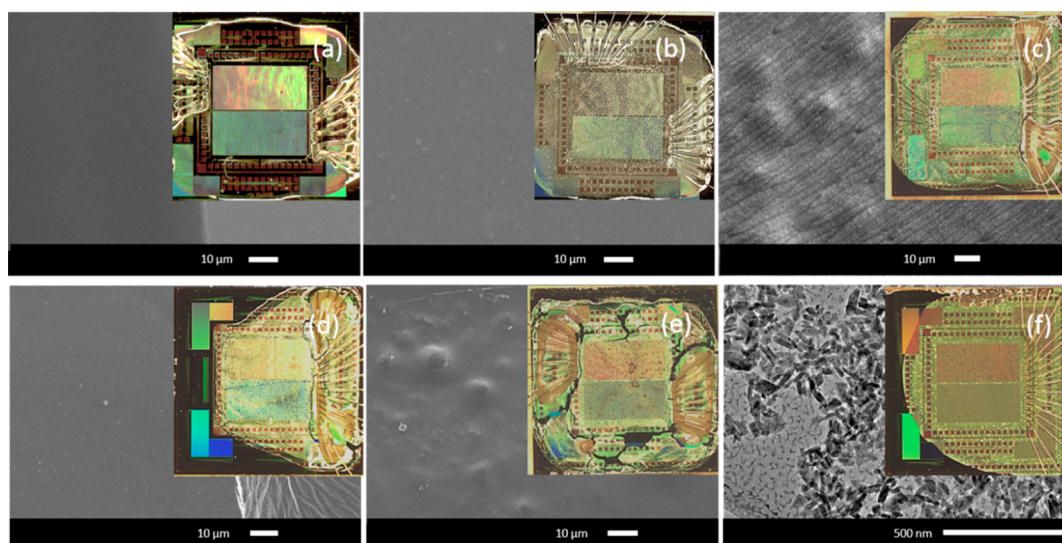
**2.4. Preparation of the Device Affinity Layers.** The polyimide (PI) Matrimid S218 was pretreated at  $453\ \text{K}$  overnight under vacuum to remove any physically adsorbed water. Different loadings ( $0$ ,  $20$ ,  $40$ ,  $50$ , and  $60\ \text{wt}\%$ ) of  $\text{NH}_2\text{-MIL-53(Al)}$  in Matrimid were prepared by dissolving the calculated amount of materials in tetrahydrofuran (THF) in such a way that the total weight of both components was  $6\ \text{wt}\%$  ( $5.5\ \text{wt}\%$  for  $50$  and  $60\ \text{wt}\%$  loadings). First, the required amount of MOF particles for each concentration was dispersed in  $940\ \text{mg}$  of THF by ultrasonication for  $15\ \text{min}$ . Then, the remaining amount of polymer (for each desired concentration) was added to this dispersion and stirred overnight. In the case of MOF-free polyimide coatings,  $60\ \text{mg}$  of PI was dissolved in  $940\ \text{mg}$  of THF and stirred overnight. For the preparation of coatings containing pure MOF ( $100\ \text{wt}\%$ , i.e., without PI),  $22.5\ \text{mg}$  of MOF nanoparticles were dispersed in  $480\ \text{mg}$  of THF. Before the deposition of the solutions over devices

and reference silicon substrates, the solutions were subjected to three alternative cycles of ultrasonication and mixing of  $15\ \text{min}$  each to ensure good dispersion of MOF particles.<sup>22,31</sup> The dispersed solution was then cast over the reference substrate and transducer devices with  $5\ \mu\text{L}$  drops under a solvent-saturated environment to reduce evaporation of the THF solvent. This ensures that the formed coatings do not crack and remain uniform. Once dried, the coated devices were treated at  $453\ \text{K}$  in vacuum for  $4\ \text{h}$ .

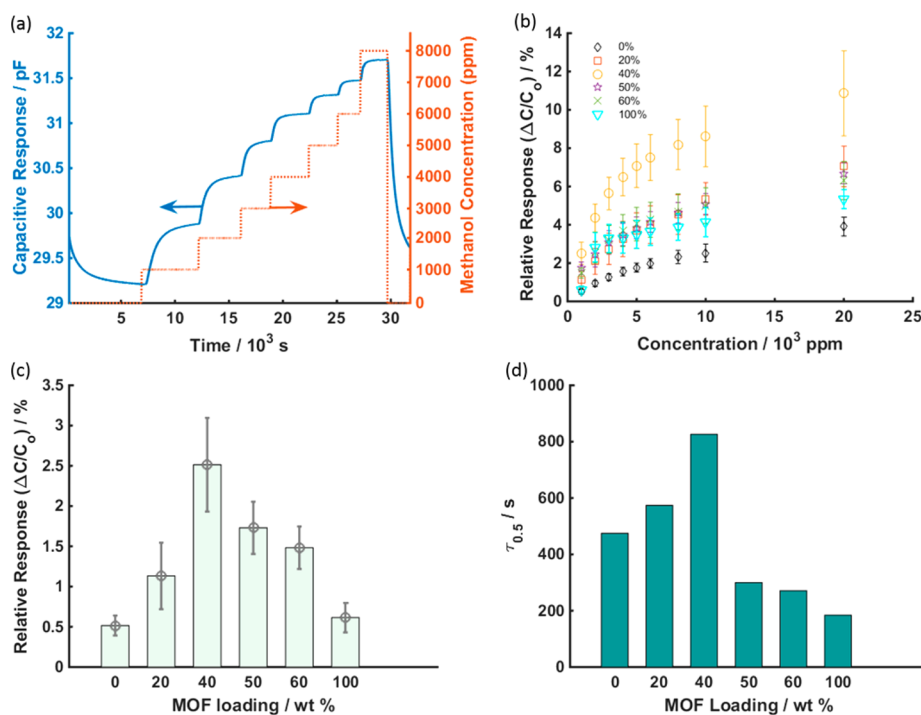
**2.5. Materials Characterization.** Scanning Electron Microscopy (SEM) images of the composite films were acquired using a JEOL JSM 6010LA microscope after sputtering the sample with a conductive gold layer. TEM images of MOF particles were acquired with a JEOL JEM-1400 instrument. XRD patterns were recorded with a Bruker D8 Advance diffractometer with  $\text{Co-K}\alpha$  radiation ( $\lambda = 1.78897\ \text{\AA}$ ). The diffraction pattern was scanned with a step size of  $0.02^\circ$  and a scan speed of  $0.2\ \text{s}$  per step. A DEKTAK 8 profilometer was used to determine the thickness of the composite films over reference silicon substrates with a force of  $0.3\ \text{mg}$  and a horizontal and vertical resolution of  $0.5\ \mu\text{m}$  and  $4\ \text{nm}$ , respectively. For  $40\text{--}60\ \text{wt}\%$   $\text{NH}_2\text{-MIL-53(Al)}$ /Matrimid devices, the thickness was measured with a Nikon MUC-151 micrometer because of the brittleness of the composite layer. Impedance measurements of the sensor devices were carried out with an HP 4284A LCR meter in a four-probe configuration.

**2.6. Gas Sensing Measurements.** An automated gas mixing and detection setup has been designed and built for monitoring the behavior of gas sensor devices using impedance measurements on exposure to analytes and vapors in the gas phase.<sup>29</sup> Figure 2 provides a schematic lay-out. The major part of the equipment consists of a (i) gas mixing unit containing mass flow controllers for providing desired flow rates of the gases, (ii) vapor generation system with a series of bubblers for analysis of individual and mixtures of vapors, (iii) temperature-controlled measurement chamber to analyze the components in the gas phase using sensor devices under test, (iv) electrical measurement units for monitoring perturbations in the measurement chamber, and (vi) instrument controller and data logger system for controlling mass flow controllers, electrical instruments and recording data using LabVIEW-based software.

The gas mixing units were connected with the calibrated gas bottles to provide gases with desired concentration and flow rates to the measurement chamber. The addition of vapors was achieved by passing the inert gas ( $\text{N}_2$ ) at atmospheric pressure through a series of bubblers to generate a saturated stream of  $\text{N}_2$  with the vapor at room temperature. The vapor concentration was calculated with the Antoine Equation:<sup>39,40</sup>



**Figure 3.** SEM micrographs and optical images (inserts) of the casted layers of  $\text{NH}_2\text{-MIL-53(Al)}$ -Matrimid composites with different MOF concentrations: (a) 0, (b) 20, (c) 40, (d) 50, and (e) 60 wt %. (f) TEM micrograph of  $\text{NH}_2\text{-MIL-53(Al)}$  nanoparticles along with the optical image of a casted device with only MOF particles (100 wt %).



**Figure 4.** (a) Quantitative, cumulative response toward methanol by a sensor device coated with a thin layer of 40 wt %  $\text{NH}_2\text{-MIL-53(Al)}$  nanoparticles in Matrimid. (b) Comparison of the sensor response by devices with different ratios of  $\text{NH}_2\text{-MIL-53(Al)}$  and Matrimid to the exposure of different concentrations of methanol. (c) Comparison of the response of these sensor devices toward 1000 ppm (0.1%) of methanol. (d) The half-time for response saturation ( $\tau_{0.5}$ ) for these sensor devices on the exposure to 20 000 ppm (2%) of methanol. The error-bars were obtained from the measurements done on three independently replicated devices.

$$\log P^{\circ} = A - \frac{B}{T + C} \quad (1)$$

where  $P^{\circ}$  is the saturated vapor pressure of the analyte, and  $T$  is the temperature of the bubbler.  $A$ ,  $B$ , and  $C$  are analyte-dependent constants.

Sensing measurements were carried out using an HP 4284A precision LCR meter at a frequency of 20 kHz and a voltage of 1 V in a four-probe arrangement. The devices were kept at a controlled temperature of 28 °C and a total flow of 200 mL/min. During the measurement, the desired concentration of an analyte component in

dry  $\text{N}_2$  was passed through the measurement chamber having a volume of  $\sim 400$  mL. After stabilization and equilibration, the sensors were recovered to the baseline by replacing the stream of  $\text{N}_2$  with the analyte by dry  $\text{N}_2$  only.

The frequency analyzer and the gas mixing setup were automated using a custom-built LabVIEW program to control the parameters such as flow rates, voltage, temperature, and carry out the time-based and frequency-based measurements.

### 3. RESULTS AND DISCUSSION

**3.1. Sensing Layer Formation.** The crystallinity of the synthesized NH<sub>2</sub>-MIL-53(Al) MOFs nanoparticles was confirmed with powder X-ray Diffraction (XRD)<sup>31</sup> (Figure S1 of the Supporting Information). The structure of the MOF consists of AlO<sub>4</sub>(OH)<sub>2</sub> octahedra coordinated to 2-amino-terephthalic acid and mainly exhibited a narrow-pore (np) configuration.<sup>31</sup> The nanoparticles had a rod-shaped morphology and dimensions of ~56 × ~17 nm<sup>2</sup> (Figures 3f and S2).

In order to investigate the role of the MOF particle concentration in the affinity layer for sensing, sensor devices were prepared with different filler ratios (0–60 wt %) of NH<sub>2</sub>-MIL-53(Al) in a matrix of Matrimid. The polymer-MOF mixture in tetrahydrofuran (THF) was drop-casted in a solvent-saturated environment onto the sensor devices and reference silicon substrates to obtain a 10 μm thick layer (Figure S3). The thickness was kept at 10 μm to ensure that the electrical field lines of the transducer stay inside the polymeric layer.<sup>29</sup> This is three times higher than the spatial wavelength (~3 μm) of the transducer and ensures that deposition variations do not influence the equilibrium sensor response. The coated region was kept larger than the central active region (Figures 1 and 3) to maintain uniformity in the active region. The differences in the deposited composite layers can be seen with SEM (Figures 3a,b) and optically (Figure 3, insets). The surface coating turned slightly opaque due to the presence of agglomerated MOF particles. The roughness of the material was increased. At further increased MOF loadings, opacity also increased and the layer became brittle. At loadings above 50 wt % the layer showed cracks (Figures 3d,e and S4). For comparison, MOF particles (100 wt %, i.e., in the absence of polymer matrix) were also deposited over the sensor devices (Figure 3f).

**3.2. Quantitative Investigation of the Sensing Performance.** Sensor devices coated with layers containing different MOF loadings were exposed to increasing concentrations of methanol in N<sub>2</sub> at 28 °C. Methanol was selected as the primary analyte due to the good sensitivity of NH<sub>2</sub>-MIL-53(Al) and Matrimid for this alcohol.<sup>29</sup> The capacitive changes on the exposure to methanol were monitored by using impedance spectroscopy with a frequency of 20 kHz and a voltage of 1 V. The sensor devices were first exposed to dry N<sub>2</sub> in the measurement chamber until a stable baseline was achieved. Next, methanol vapors were introduced at different concentrations using a series of two bubblers (see Section 2.6 for details). Figure 4a shows a typical response of a sensor device covered with a PI coating containing 40 wt % of MOF nanoparticles toward different concentrations of methanol at 28 °C. The methanol concentration was kept between 1000 and 20 000 ppm (0.1–2%), which is in line with industrial requirements.<sup>41,42</sup> Lower concentrations (below 100 ppm) were not reliable due to the equipment limitations.

The capacitive response of the sensor device increases on exposure to methanol in the measurement chamber ( $t = \sim 7000$  s in Figure 4a). The capacitive response further increased upon increasing methanol concentrations. Furthermore, the response was completely reversible as the initial baseline was recovered on replacing the methanol containing N<sub>2</sub> gas by dry N<sub>2</sub> (Figures 4a and S5).

The sensitivity of our devices with different amounts of MOF in the polymer matrix toward methanol is plotted in Figure 4b. The response was normalized by  $C_0$ , the capacitance in dry N<sub>2</sub>, yielding a relative change in capacitance ( $\Delta C/C_0$ ) allowing easy

and accurate comparison of the different sensor devices. The inclusion of these porous nanoparticles inside the polymer matrix enhances the overall signal response, e.g., the addition of 20 wt % of MOF more than doubles the sensor response to methanol. This increase in the sensitivity is due to the additional adsorption sites introduced by the NH<sub>2</sub>-MIL-53(Al) along with the intrinsic adsorption capacity of the Matrimid.<sup>29,38,43</sup> The response increased further upon increasing amounts of MOF in the polymeric layer, resulting in a 5-fold increase (over exposure to 1000 ppm methanol) for the 40 wt % MOF loading as compared to the MOF-free Matrimid layer (Figure 4b,c). Further, the maximum capacitive response was observed at 40 wt % (Figure 4c). Such a response (~750 fF of capacitance change) is clearly higher than the detection limit of the instrument (0.1 fF), making the current approach suitable for measuring methanol within most industrial relevant concentrations (0–3000 ppm) at ambient conditions.<sup>41</sup> It also allows to compete well with the commercially available methanol sensors (having a measurement range within 0–10 000 ppm). The observed decrease in the response at a higher wt % indicates a transition from a polymer-limited adsorption to a MOF-limited adsorption (Figure 4b,c).

We attribute this reduction in response to the decreased adhesion between the composite layers and transducer at loadings above 40 wt %. Formation of brittle and cracked films at these loadings, as observed by SEM, supports this line of reasoning (Figure 3). Additionally, this decreased signal transduction behavior is also clearly visible in the sensor response of 100 wt % MOF devices as the high adsorption capacity of NH<sub>2</sub>-MIL-53(Al) is not completely reflected in the sensor response (Figures 4b and S6). Moreover, percolation of methanol molecules from MOF particle to MOF particle at such high loadings (above 40 wt %) ensures that the adsorption process is directly determined by the MOF.<sup>23</sup> The slightly higher response in these devices compared to that for the 100 wt % MOF-coated device (Figure 4b,c) can be explained by additional adsorption in the Matrimid phase, which covers these nanoparticles partially.

The response kinetics provide further insights into the adsorption process. Interestingly, considering the time taken to reach half the saturation response ( $\tau_{0.5}$ ), the  $\tau_{0.5}$  is observed to increase (slower process) with increasing the amount of NH<sub>2</sub>-MIL-53(Al) in the Matrimid matrix for the lower concentration regime (0–40 wt %) (Figure 4d). The observed increase of  $\tau_{0.5}$  is attributed to rigidification of the polymer matrix near the polymer-MOF interface.<sup>44,45</sup> The solvent evaporation can induce stress in the polymeric matrix around the filler molecules and reduce the relaxation of the polymeric chains close to the MOF surface while drying takes place. This would lead to a decrease in the free volume close to the surface and hence reduction in the permeability. This has also been observed before in membranes prepared by spin-coating layers of a glassy polymer (PSF-Udel) with NH<sub>2</sub>-MIL-53(Al).<sup>45</sup> In addition, partial penetration of the Matrimid polymer chains into the micropores of NH<sub>2</sub>-MIL-53(Al) can also affect the transport of methanol to the MOF particles.<sup>45</sup> Also, as shown in Figure 4d,  $\tau_{0.5}$  drops by a factor of approximately 3 (faster process) at a MOF loading of  $\geq 50$  wt % when compared with devices containing 40 wt % of MOF. At loadings of 50 and 60 wt % the response times also became significantly closer to the one of the MOF-coated device (100 wt %). This observation confirms the direct accessibility of the MOF particles to the exposed environment due to cracks and to a thin or partial

coverage of NH<sub>2</sub>-MIL-53(Al) particles with the (rigidified) Matrimid.

Moreover, on the assumption that the sensor response is proportional to the adsorbed concentration of the analyte in the composite matrix,<sup>29</sup> the nonlinear response of different devices with varying methanol concentrations (Figure 4b) was fitted to the Langmuir adsorption model as shown in eq 2:

$$\nu = \nu_{\max} \times \frac{K_{\text{eq}} \times p_{\text{MeOH}}}{1 + K_{\text{eq}} \times p_{\text{MeOH}}} \quad (2)$$

where  $\nu$ ,  $\nu_{\max}$ ,  $K_{\text{eq}}$ , and  $p_{\text{MeOH}}$  indicate the observed response ( $\Delta C/C_0$ ), maximal (saturation) response ( $\Delta C_{\text{max}}/C_0$ ), adsorption equilibrium constant, and partial pressure of the analyte, respectively.  $\nu$  and  $\nu_{\max}$  are represented as normalized response (%) for accurate comparison between different devices as the  $C_0$  varies in each device due to the differences in the dielectric properties of composite films. Since the concentration range was still in the lower regime of the adsorption isotherm, we assumed that the adsorption behavior is still following the Langmuir model. The fitting was carried out on the average response by devices with different MOF loadings in the polymer matrix (Figure S7). The obtained parameters are listed in the Table 1, showing that there is good agreement with the experimental data in all the polymer-MOF coatings ( $R^2 \geq 0.978$ ).

**Table 1. Langmuir Parameters ( $K_{\text{eq}}$  and  $\nu_{\max}$  ( $=\Delta C_{\text{max}}/C_0$ )) for Sensor Devices Coated with a Matrimid Layer Containing Different Amounts of NH<sub>2</sub>-MIL-53(Al) as Function of Different Vapor Concentrations of Methanol**

MOF loading (wt %)	$K_{\text{eq}}$ (bar <sup>-1</sup> )	$\nu_{\max}$ (%)	$R^2$	$\nu_{\max, \text{MOF}}$ calculated (%)
0	120	4.8	0.996	0
20	132	8.9	0.999	25
40	241	12.9	0.995	25
50	368	6.1	0.978	7.4
60	264	7.1	0.998	8.6
100 <sup>a</sup>	860	4.4	0.947	4.4

<sup>a</sup>Sensor devices with 100 wt % MOF showed deviation from Langmuir behavior at low concentrations (1000–2000 ppm (0.1–0.2%)) of methanol.

The 100 wt % coating devices showed a slight deviation from Langmuir behavior at low methanol concentrations. This can be attributed to additional structural changes in the NH<sub>2</sub>-MIL-53(Al) due to the initial adsorption of methanol in freely

accessible MOF,<sup>38</sup> which resulted in a significant transition in the response between 1000 (0.1%) and 2000 ppm (0.2%) of methanol in devices coated with 100 wt % MOF (Figure 4b).

Further, the equilibrium affinity constants of experiments using devices with NH<sub>2</sub>-MIL-53(Al) (100%) and a Matrimid coating show that both MOF and polymer have an intrinsic affinity toward methanol. MOF particles showed higher affinity as expected due to the hydrophilic nature of the pores and the high porosity. All composite-coated devices showed the affinity constants in-between these two extreme cases with 40 wt % showing the highest saturation response with an increase of 12.9%, compared to 4.8% and 4.4% for the only Matrimid (0 wt %) and only-MOF (100 wt %) coated devices, respectively.

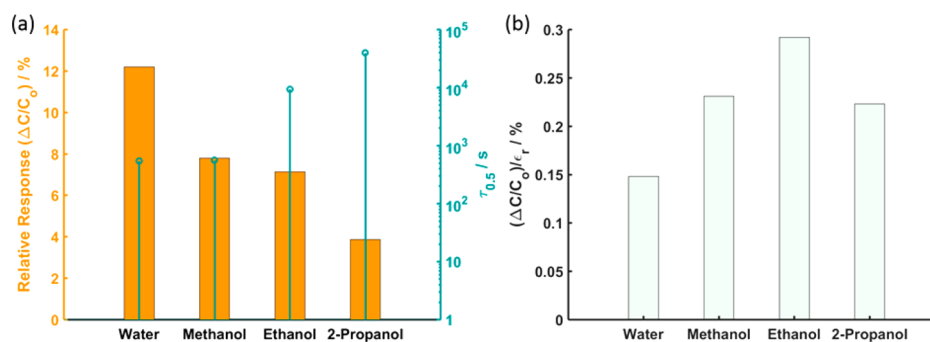
On the basis of these saturation responses and the mass fraction of MOF, the contribution of MOF toward the response ( $\nu_{\max, \text{MOF}}$ ) was calculated in these composite films using eq 3:

$$\nu_{\max} = \nu_{\max, \text{MOF}} \times w_{\text{MOF}} + \nu_{\max, \text{Matrimid}} \times w_{\text{Matrimid}} \quad (3)$$

where  $\nu$  and  $w$  are the equilibrium response and the mass fraction of MOF or Matrimid, respectively. It was assumed that the adsorption contributions of both Matrimid and NH<sub>2</sub>-MIL-53(Al) are additive in the total response. The contribution of Matrimid was determined as the mass fraction of the MOF-free Matrimid coated device. Expectedly, the calculated contribution of MOF toward the relative response doubled over increase in loading from 20 wt % ( $\nu_{\max, \text{MOF}} \times w_{\text{MOF}} = 5.0\%$ ) to 40 wt % ( $\nu_{\max, \text{MOF}} \times w_{\text{MOF}} = 10.0\%$ ). From both cases, the saturation response ( $\nu_{\max, \text{MOF}}$ ) of the MOF-only (100 wt %) device was calculated to be 25% which is nearly 6 times higher than the observed response (4.4%). Such decrease in  $\nu_{\max}$  (also in case of devices with 50 and 60 wt % MOF (Table 1)) can be associated, as discussed above, to the decreased adhesion in these highly loaded devices.

In summary, from this comparative study it can be concluded that 40 wt % showed the highest capacitive change and that for increasing MOF loadings there was a transition at 40 wt % from polymer-driven adsorption to a MOF-driven adsorption.

**3.3. Sensitivity and Selectivity.** Next to sensor responses and kinetics, another important parameter to consider for a sensor device is its cross-sensitivity toward different analytes. The 40 wt % MOF-modified sensor devices were exposed to water, methanol, ethanol, and 2-propanol to study its cross-sensitivity (Figures 5 and S8). Upon exposure to 20 000 ppm (2%) of each analyte, the equilibrium response decreased and  $\tau_{0.5}$  increased with increase of molecular size of the analyte (water, methanol, ethanol, and 2-propanol) (Figure 5a). A 10-fold and 70-fold increase in  $\tau_{0.5}$  for ethanol and 2-propanol,



**Figure 5.** (a) Relative response of a sensor device containing 40 wt % MOF in a Matrimid layer upon the exposure to 20 000 ppm (2%) of different analytes (orange) and the corresponding half-time for saturation,  $\tau_{0.5}$  (blue). (b) Relative response normalized to the relative dielectric constant.

respectively, compared to methanol and water clearly indicates slower uptakes, likely due to a slower diffusion through the Matrimid polymer composite.<sup>29</sup> In order to correlate with the analyte concentration, the equilibrium responses were normalized to the relative dielectric constant ( $\epsilon_{r,\text{Water}} = 78.3$ ,  $\epsilon_{r,\text{Methanol}} = 32.7$ ,  $\epsilon_{r,\text{Ethanol}} = 24.5$ , and  $\epsilon_{r,\text{Propanol}} = 18$ , Figure 5b). Now, it can be seen that the uptake of analytes by the MOF-Matrimid composite sensor increases from water to methanol to ethanol. This is well in line with the known higher affinity of Matrimid toward ethanol and methanol than water.<sup>46,47</sup> Even though the affinity of Matrimid toward 2-propanol was also higher,<sup>46,47</sup> it still resulted in a decrease in the normalized response. This decrease along with the slower uptake can be attributed to the size-exclusion effect of the Matrimid as previously observed, with no response toward 2-propanol in only Matrimid based sensors.<sup>29</sup> In contrast, the presence of a higher amount of MOF in the polymer matrix allowed some adsorption of this analyte.

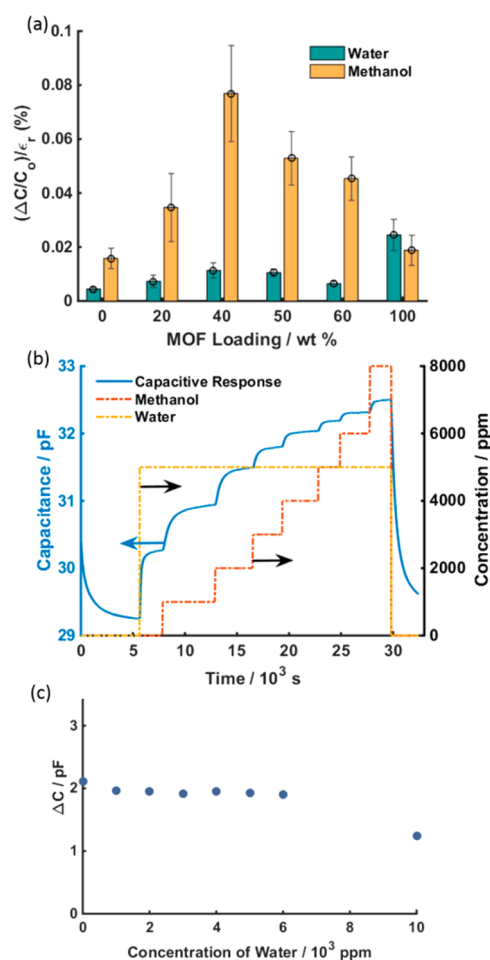
Moreover, with significant responses and similar  $\tau_{0.5}$  toward water and methanol, these sensor devices show sensitivity toward both analytes. In order to study this cross-sensitivity further, the behavior of devices with different MOF loadings was compared on exposure to 1000 ppm of water and methanol at 28 °C (Figure 6a). The change in the response was normalized with respect to the relative dielectric constant of the analyte to make interpretation in terms of change of concentrations in the affinity layer possible.

Figure 6a shows that sensors with only Matrimid as affinity layer exhibit a higher response toward methanol than to water. This is due to the intrinsic higher affinity of Matrimid toward methanol compared to water.<sup>29,46</sup> For sensors having an affinity layer of pure MOF, the response toward water is slightly larger than toward methanol. This is attributed to the slightly hydrophilic nature of the MOF (certainly the MOF is more hydrophilic than the polymer). For the MOF-Matrimid composite systems, the uptake of water and methanol is always higher than for the pure Matrimid system, with a maximum response (uptake) observed around 40 wt % MOF.

In Figure 6b the sensor response to varying methanol concentrations in the range of 1000 to 8000 ppm in steps of 1000 ppm in the presence of a fixed water concentration of 5000 ppm is presented. The sensor has an affinity layer containing 40 wt % MOF in Matrimid. The response is similar (after correction for the fixed contribution of water to the capacitive response) to the response observed in the absence of water (Figure 4a), indicating that there is no competition for water and methanol affinity, and both analytes do behave as if the other component is not present.

In Figure 6c the sensor response to a fixed concentration of methanol of 5000 ppm (0.5%) is shown in the presence of varying water concentrations (0–10 000 ppm (0–1%)). In the investigated range until 6000 ppm (0.6%) of water, it is observed that the capacitance increase is independent of the water concentration, confirming that methanol and water adsorption is noncompetitive. Deviation was observed at still higher water concentration, where condensation in the pores is likely to occur.

**3.4. Effect of Temperature on Sensor Response.** To get further insight into the sensing process, adsorption studies of methanol and water were also carried out at different temperatures. A sensor device with 40 wt % NH<sub>2</sub>-MIL-53(Al) in Matrimid and a MOF-free Matrimid (0 wt %) were exposed to 1000 ppm (0.1%) of methanol and 1000 ppm



**Figure 6.** (a) Relative response normalized to relative dielectric constant of sensor devices containing different loadings of NH<sub>2</sub>-MIL-53(Al) in a Matrimid layer upon the exposure to 1000 ppm (0.1%) of water and methanol. The error-bars were obtained from the measurements performed on 3 independently replicated devices. (b) The sensor response of a device with 40 wt % NH<sub>2</sub>-MIL-53(Al) in Matrimid toward different concentrations (1000 to 8000 ppm (0.1 to 0.8%) in steps of 1000 ppm (0.1%), indicated by right axis) of methanol in the presence of 5000 ppm (0.5%) of water. (c) The capacitance change of a device with 40 wt % NH<sub>2</sub>-MIL-53(Al) in Matrimid toward exposure of 5000 ppm (0.5%) of methanol. Each measurement point was obtained at different initial water concentrations as indicated by the x-axis. The temperature was kept at 28 °C.

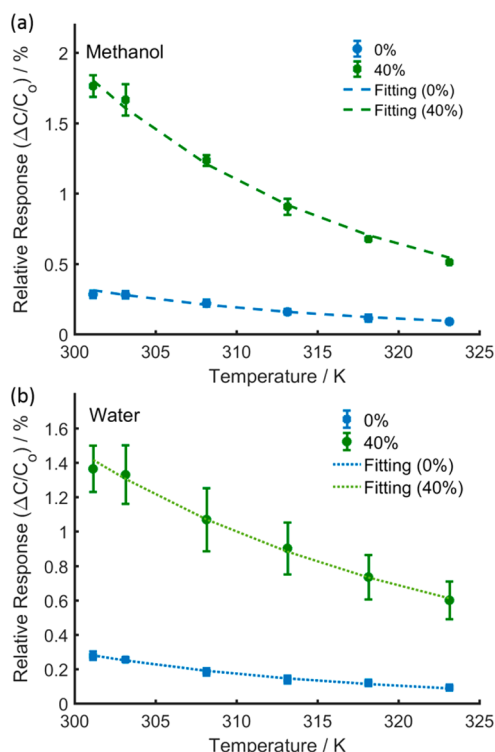
(0.1%) of water at different temperatures (301–323 K). The effect of the operating temperature on sensing performance is clearly visible in Figures 7 and S9, showing an Arrhenius plot of the relative response ( $\nu$ ) (defined as  $\Delta C/C_0$ ) versus the absolute temperature ( $T$ ). At higher temperatures, the capacitance change  $\Delta C$  decreases, indicating that less analyte is adsorbed by the affinity layer, as expected.

For the classical Arrhenius behavior applies:

$$\nu = \nu_0 e^{-\frac{\Delta H}{RT}} \quad (4)$$

where  $\nu_0$  and  $R$  indicate a pre-exponential constant and the gas constant, respectively.  $\Delta H$  represents the difference in activation energy of adsorption and desorption and  $\nu$  represents the relative change in the capacitance ( $\Delta C/C_0$ ). From the exponential fit of the Arrhenius plot (Figure 7), we have deduced the value of  $\Delta H$ .





**Figure 7.** Arrhenius relation between the relative response ( $\nu = \Delta C/C_0$ ) and the temperature ( $T$ ) for devices coated with 40 wt %  $\text{NH}_2\text{-MIL-53(Al)}$  in Matrimid (green) and Matrimid only (blue) on the exposure to 1000 ppm (0.1%) of methanol (a) and 1000 ppm (0.1%) of water (b) at different temperatures. The error-bars were obtained from three measurements done on the same device.

The enthalpy of adsorption ( $\Delta H$ ) for devices with only Matrimid-coated devices was calculated to be  $-44.4 \pm 3.0$  kJ/mol and  $-42.6 \pm 9.2$  kJ/mol for methanol and water, respectively. A negative value of  $\Delta H$  confirms the exothermic nature of the process, with similar affinities for methanol and water. For devices with 40 wt % MOF loading, the enthalpy of adsorption ( $\Delta H$ ) was found to be  $-44.1 \pm 1.3$  kJ/mol and  $-30.9 \pm 1.5$  kJ/mol for methanol and water, respectively, clearly indicating higher affinity for methanol than for water. Moreover,  $\Delta H$  has become less favorable by the addition of MOF to the polymer for adsorption of water. We tentatively suggest that these differences in  $\Delta H$  can be related to polymer rigidification near the MOF interface.<sup>44,45</sup> This rigidification would result in a stronger barrier toward water than toward (the less polar) methanol.

#### 4. CONCLUSIONS

In this study, we have investigated affinity layers composed of varying amounts of  $\text{NH}_2\text{-MIL-53(Al)}$  MOF mixed in a Matrimid polymer and deposited on top of meander-patterned aluminum electrodes for capacitive sensing in the presence of different analytes in the gas phase. As analytes, we have investigated water, methanol, ethanol, and 2-propanol. The capacitive response toward methanol and water was enhanced by the incorporation of the MOFs in the polymeric matrix. The optimum composition was found at  $\sim 40$  wt % of MOF, showing an up to 5-fold increase in equilibrium response for 1000 ppm of methanol compared with only Matrimid coated devices. The response time also increased with increasing MOF concentration (up-to 40 wt %), which is attributed to polymer

rigidification near the MOF-polymer interface. Devices containing  $\geq 50$  wt % MOF in Matrimid resulted in brittle and cracked films. The equilibrium response and response time decreased in these devices compared to devices with 40 wt % MOF and became closer to the equilibrium response and response time shown by a polymer-free, MOF-coated sensor device. The decrease in the response behavior was attributed to the poor adhesion. Quantitative analysis of the sensor devices indicated Langmuir behavior showing a maximum affinity ( $K_{\text{eq}} = 860 \text{ bar}^{-1}$ ) for the polymer-free, MOF-coated devices. The affinity of all composite devices was found to be in-between those of 100% MOF and 100% Matrimid-coated devices, with the 40 wt % devices showing a maximum saturation response. Cross-sensitivity studies showed the ability of the devices to kinetically distinguish between alcohols and water. This is based on the molecular size of the analyte resulting in a  $\sim 70$  times increased  $\tau_{0.5}$  for 2-propanol for the 40 wt %-MOF-polymer sensor compared to methanol and water. Sensing studies with mixtures of water and methanol indicated that both molecules have independent contributions. For water this additive effect to the detection of methanol holds up to concentrations of 6000 ppm. The temperature dependency of the sensor devices showed an Arrhenius-type behavior with a stronger methanol adsorption than water adsorption, confirming a higher affinity toward methanol. With differences in response sensitivity and kinetics for devices coated with affinity layers with different MOF-polymer ratios, quantitative fingerprinting of various analytes is within reach by utilizing sensor arrays with different MOF-polymer fractions and types at different operational temperatures.

#### ■ ASSOCIATED CONTENT

##### Supporting Information

The Supporting Information is available free of charge on the ACS Publications website at DOI: 10.1021/acsami.7b02630.

XRD and TEM data of  $\text{NH}_2\text{-MIL-53(Al)}$  nanoparticles (Figures S1 and S2), thickness profiles, SEM data, capacitive sensor response, methanol adsorption isotherms, concentration-dependent sensor response (and Langmuir fittings), cross-sensitivity data, and temperature-dependent sensor response of devices with different composite coatings (Figures S3–S9) (PDF)

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##### Notes

The authors declare no competing financial interest.

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